

the inch, it is best to proceed by direct percolation, after moistening the powder with a fourth of its weight of alcohol.

Fleming's Tincture of Aconite Root should be made with the same precaution; and in this case it is even more necessary, as the proportion of root to tincture is larger.

Aconite Liniment is the name given to a concentrated solution of aconite, with a portion of glycerin, originally prepared by the writer of this paper. An experience of eight years with this preparation, warrants the opinion that it is strongly deserving of the attention of the medical profession. A fluid ounce of this preparation represents two ounces of the root; it is therefore about five times the strength of the Official Tincture, and about three times that of Fleming's Tincture. I have for years past adopted the custom of keeping a normal solution of aconite root, obtained by thoroughly exhausting the drug with the alcohol, and evaporating the tincture until each fluid ounce represents two ounces of the root, removing the fixed oil that separates, and if necessary agitating it with a little diluted alcohol, to remove any adherent aconitia and extractive matter. From this the weaker preparations are made by dilution, and the stronger by evaporation, thus:—

Tinctura Aconiti Radicis. U. S. P.—Take of normal solution of aconite, three fluid ounces; alcohol, twelve fluid ounces. Mix them.

Fleming's Tincture of Aconite Root.—Take of normal solution of aconite, five fluid ounces alcohol, ten fluid ounces. Mix them.

Aconite Liniment.—Take the normal solution of aconite, half a pint; glycerin, a fluid ounce. Evaporate the solution to seven fluid ounces, add the glycerin, and mix them.

Aconite Plaster.—Take of normal solution of aconite root, half a pint; adhesive plaster, a sufficient quantity. Evaporate the solution of aconite until it assumes the consistence of thick honey, then, having weighed it, add as much adhesive plaster, previously melted, as will make the whole weigh sixteen ounces.

Extractum Aconiti Radicis.—Take of normal solution of aconite, half a pint. Evaporate it carefully on a water bath until reduced to the proper consistence for an extract.

In regard to the merit of this mode proceeding, it can be said that it assures the thorough extraction of the aconite, and promotes uniformity in the strength of the preparations, and also affords a very great convenience to the dispenser, who can by preparing a quantity of the strong solution quickly produce the other preparations at will, in quantities to suit his requirements.—*Chemist & Druggist.*

A NEW SALT OF IRON AND QUININE.

By DR. FERGUS, of Marlborough College.

It is generally found that a salt of the protoxide of iron is preferable to one of a higher degree of oxygenation; but it is also difficult to obtain an absolutely permanent salt of the protoxide. Perhaps without exception, the sulphate is the most practically useful of all the salts of iron, owing to the uniformity of its composition. Of the quinine salts, the sulphate is also the most available for general purposes. It is not difficult to form a simple combination of these two sulphates, but the resulting compound is not well fitted for general use. The addition, however, of a certain proportion of sulphate of magnesia, enables us to obtain a salt which is nearly as soluble as the sulphate of magnesia itself—quite unalterable in the solid state, and forming a solution perfectly clear at first, and remaining so for an indefinite period. The iron has no tendency to a further state of oxygenation; the solution had been agitated with oxygen gas, and kept in contact with it for several days, without the least change. A solution of gallic acid tinges a solution of the salt of light bluish color after the lapse of two or three days, and many substances which produce an inky compound with the salts of iron may be mixed with it without causing any change of color.

The proportion of the three sulphates which has been adopted, is 80 per cent. of sul-