

It appears from this tabulation that the mean strength of vinegar as sold in Canada in 1905, is distinctly higher than in 1893. The change is, however, more marked if comparison be made as in the following table:—

District.	Below 3 p. c.	From 3 to 4 p. c.	From 4 to 5 p. c.	From 5 to 6 p. c.	From 6 to 7 p. c.	Above 7 p. c.	Remarks.
1893.							
Nova Scotia.....	1	10	6	1	1	0	p. c.
Quebec.....	1	1	0	11	7	3	7·41, 7·80, 11 14.
Ontario.....	0	2	14	16	5	2	7·99, 7·27.
Manitoba.....	0	3	3	4	0	2	7·96, 7·42.
1905.							
Nova Scotia.....	2	7	6	3	1	1	8·30.
New Brunswick.....	1	5	6	1	5	2	7·50, 7·04.
Prince Edward Island	0	7	4	5	1	3	7·98, 9·72, 7·07.
Quebec.....	0	0	0	6	9	5	9·04, 7·40, 9·15, 9·26, 7·07.
St. Hyacinthe.....	0	0	6	3	8	3	7·20, 7·92, 7·30.
Montreal.....	0	0	4	9	6	1	7·74.
Province of Quebec.	0	0	10	18	23	9	
Kingston.....	0	0	10	7	2	1	7·68.
Toronto.....	2	1	8	8	1	0	
London.....	0	2	9	7	1	1	9·20.
Province of Ontario	2	3	27	22	4	2	
Manitoba.....	0	0	8	3	3	6	12·60, 10·80, 10·92, 8·88, 7·82, 7·60.
Calgary.....	1	1	2	6	1	9	9·69, 8·94, 10·38, 9·93, 11·04, 10·56, 11·97, 7·68, 9·48.
British Columbia....	1	1	8	4	0	8	9·78, 9·71, 9·30, 10·86, 9·77, 9·06, 9·00, 10·56.

A study of this table shows that, by provinces, the percentage of samples of vinegar containing more than 7 per cent of acetic acid, is as follows:—

Nova Scotia.....	5 p. c.	Ontario.....	3 p. c.
New Brunswick.....	10 p. c.	Manitoba.....	30 p. c.
Prince Edward Island	15 p. c.	North-west Territories	45 p. c.
Quebec.....	15 p. c.	British Columbia....	37 p. c.

It is apparent that by far the larger number of highly acid vinegars are sold in the provinces of Manitoba, British Columbia and the North-west Territories. It is known that a substance sold as concentrated vinegar, or some equivalent name, finds a market in the west. A sample of this article (see file 90457) was examined in this laboratory in May of this year, and was found to contain 55·2 per cent of acetic acid. In reporting on this article Mr. Macfarlane says: 'I am of opinion that this sample is a preparation from concentrated acetic acid, and that it would give, on dilution, a factitious vinegar, which it would be impossible to distinguish from the genuine article.' It seems to me quite probable that the samples giving such abnormally high acid strength are made from this or similar concentrates, if not, then acetic acid has been added to a normal vinegar in order to give it the acid value found.