without disturbing the precipitate any more than possible; then add a little hot water; stir up with rod. Gradually wash out all the precipitate with hot water. Filtrate must be quite bright. Wash carefully the filter and precipitate with hot water until a drop from funnel gives no reaction with silver nitrate. Dry filter in funnel in water oven.

Transfer to the crucible as much as possible of the precipitate, using platinum wire. Fold the dry filter tightly lengthwise with remaining precipitate inside, perforate top, and attach to platinum wire as instructed. Burn it over crucible as instructed; crucible should rest on glazed paper to collect any particles which might escape.

After ashing filter cover crucible, add 1 drop of dilute sulphuric acid, and heat to red heat for half an hour or until ash is quite white. Cool in desiccator, and weigh.

Calculate weight of sulphate ions; Barium sulphate multiplied by 0.4206 = H_2SO_4 by 0.412 = SO_4 ions.

SECOND LABORATORY PERIOD

URINE—NEGATIVE IONS (*continued*)

B. *Inorganic Sulphates* (Folin).—Place 25 c.c. of urine and 100 c.c. of water in a 250 c.c. Erlenmeyer flask. Acidify with 10 c.c. of hydrochloric acid (1 volume of concentrated HCl to 4 volumes of water). Add 10 c.c. of 5 per cent. barium chloride, drop by drop, as in the previous exercise, and proceed as there directed.