

solution be at the boiling point, the precipitate is amorphous; but, if the temperature be reduced to 180° , then beautiful crystals are formed, rather larger than those of sulphate of Quinine. The crystals are needle-shaped, and when magnified appear four times as large in proportion as when the solutions are cold. When they are obtained by the spontaneous evaporation of the alcoholic solution, they appear like broken fragments of cubes: the different crystalline layers remotely resemble the same appearance which is found in Ferrocyanide of Potassium. They are right-angled and well-formed: some appear tabulated. I also examined a specimen prepared by boiling the freshly precipitated crystals in water. To the eye, it looks like an amorphous powder; but, by the microscope, as well defined four-sided crystals, with straight edges. The solid angles are perfect, except in some places where they were replaced by planes; being parallelepipeds, they will belong to the square prismatic system, two of the axes being equal and the other at right angles. This salt is odourless, and very bulky. It is much less soluble than the pure alkaloid, requiring between seven and eight thousand parts of water for its complete solution, which it renders very bitter. It is much more soluble in boiling water, two drachms of which dissolve nearly $\frac{1}{3}$ ths of a grain, and retain it when cold; thus being four or five times as soluble in hot water as the free alkaloid, and not so soluble when the fluid is cold. Its behaviour with alcohol is more remarkable. When this fluid is cold, $\frac{1}{3}$ ths of a grain are soluble in two drachms, or $1\frac{1}{2}$ grains in an ounce, whereas only one grain is dissolved in the ounce when boiling; and, if a saturated cold solution be boiled, the excess of the cyanide is precipitated, and this again dissolves as the alcohol cools. When water is added to the cold solution, a certain amount is precipitated. It is soluble in all the diluted mineral acids, and in acetic acid; a small proportion also is dissolved by Hydrocyanic acid. From these solutions it is precipitated as a crystalline powder by alkalies. It is also soluble in the neutral salts, as acetate and sulphate of ammonia and potassa, and apparently also in their free alkalies. Ether dissolves a small amount. It is not soluble in Iodide, or Cyanide of Potassium, or in turpentine, at a boiling heat. Nitrate of silver gives no evidence of cyanogen in any of its solutions; but I afterwards obtained evidence of its presence by adding an excess of caustic potassa to its acetic acid solution, and then the aceto-nitrate of silver, (prepared by adding one-third of the bulk of acetic acid to a solution of the nitrate, containing a drachm of the salt to an ounce of water), by adding an excess of strong nitric acid to the precipitate, the cyanide separated in a white curdy form. I also obtained this evidence by adding the aceto-nitrate to an alcoholic solu-

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