mixture can be taken to be exactly the same as the value found by observations on the simple solutions only in the case of infinitely dilute mixtures.

The fact that for mixtures of nearly saturated solutions of these salts, the difference between the calcu ated and observed values is only 1.47 per cent, while for solutions of KCl and NaCl as near saturation, the difference is about 5 per cent, as determined by MacGregor* and again by McIntosh*, would seem to indicate that in the case of this class of salts, the magnitude of the differences depends on the amount of salt in the solution, not on the nearness to saturation.

It would appear from the above results that for mixtures of solutions of these salts not more concentrated than 0.8 equivalent gramme-molecules per litre, it is possible, by the aid of the dissociation theory, to calculate the conductivity within, or but little beyond, the limit of the error of observation.

*Loc. cit.