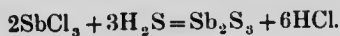


At the close of the paper an experiment* was shown, illustrating the reversibility of the reaction,



An antimony solution was prepared by dissolving 2 grammes of tartar-emetie in 20 cc. of hydrochloric acid sp. gr. 1.148, and diluting with 80 cc. of water.

- (a) Passed a little H_2S into 5 cc. antimony solution - \longrightarrow
- (b) Added 15 cc. hydrochloric acid (sp. gr. 1.148) - - \longleftarrow
- (c) Added 5 cc. antimony solution - - - - - \longrightarrow
- (d) Heated over a Bunsen burner - - - - - \longleftarrow
- (e) Cooled again in a dish of water - - - - - \longrightarrow
- (f) Added 10 cc. hydrochloric acid - - - - - \longleftarrow
- (g) Saturated with H_2S under 4 atmospheres pressure \longrightarrow
- (h) Reduced the pressure to 50 mm. (by a filter pump) \longleftarrow

NOTE.—The arrows indicate the direction of the reaction; \longrightarrow meaning precipitation of Sb_2S_3 , and \longleftarrow formation of antimony chloride from the sulphide and hydrochloric acid.

* W. Lash Miller and F. B. Kenrick, Lecture Experiments, Reversible Chemical Reactions. *Jour. Amer. Chem. Soc.* xxii. 291 (1900).