

Production and Market for Sulphuric Acid.

The gases leaving the roasting furnaces are cleaned and the sulphur dioxide they contain is converted into sulphuric acid, probably by the "contact" process. For this process it is necessary to clean the gases very perfectly, so that all dust and smoke or fume is completely eliminated, and after passing through the sulphuric acid plant, nothing will escape into the air but a proportion of the sulphur dioxide, probably less than ten percent, which has escaped conversion into acid.

Theoretically, one ton of sulphur will produce three tons of 100 percent sulphuric acid, and in view of the probable amount and analysis of the concentrates and the probable percentage recovery of the acid, it appears that a plant producing one hundred tons of zinc per day should also produce about two hundred tons of sulphuric acid.

I understand that the sulphuric acid used in and near Montreal amounts to some 10,000 tons per annum, or less than 30 tons a day, so that there would be no local market at present even for one hundred tons of additional acid, and it would probably be necessary to develop some auxiliary industry, such as the production of fertilizer, in order to use the greater part of the acid that would be produced. An ample supply of cheap sulphuric acid would tend to attract other chemical industries that use large quantities of this acid. In the meantime, the difficulty of storing and shipping sulphuric acid may tend to limit the expansion of the zinc plant.