

BLANK'S  
FAMOUS.

## Toilet Soap

6 CENTS  
PER CAKE

This is the usual 10 cent cake. I bought a big lot of it at big lot prices. I knew it wouldn't go out of style. Its better to sell soap on small profits than to sit around complaining of the hard times.

Now is your Chance.

**RICHARD ROE, Druggist,**

10 Hustle Ave. - PILLVILLE.

if they don't hurry. Everybody knows that he keeps the same kind of goods as his neighbor, Doe, does, and compounds just as carefully, but they didn't know he was selling that soap at 6 cents per cake. They will go for the soap, but at the same time they cannot help seeing his entire stock of goods, and quite likely when they leave home to get some of that soap, they will conclude that as they are going to Richard's place of business they might as well buy a tooth brush, as they are saving a few cents on soap and can thus afford it.—*Pharm. Era.*

### OLEATES, P. B.

By EDWIN WILLIAMS.

Read at a meeting of the Liverpool Chemists' Association.

We have two official oleates and a number of oleo-palmitates, such as lead plaster, ammonia liniment, carron oil, hard soap, and oleo-stearate as curd soap, each and all of indefinite composition. I will confine my remarks to what are classed as oleates in the B. P. The formula for real oleates is  $M.C._2H_{33}O_2$  or  $M''(C_{18}H_{33}O_2)_2$  according to the quantivalence of the metal. There are likewise acid oleates. The oleates dissolve in cold absolute alcohol, ether, and petroleum benzine, a property by which they may be distinguished and separated from the palmitates and stearates.

Oleates were expected to replace in the Pharmacopœia a great number of inert ointments now in use. A medicament dissolved in one will prove more efficacious by penetrating deeper into the tissues than an insoluble powder distributed on the surface. That the oleates have not found that important place anticipated for them is possibly largely attributed to the fact that our oleates are not worthy of the title, being simply a solution of the metal or oxide in oleic acid. They are of indefinite strength and most of the samples that I have obtained contained palmitate or stearate, derived perhaps from the impure

oleic acid, although some samples contained sufficient of these impurities to warrant my saying, that they were derived from the soap that they had been precipitated from.

Oleate of mercury is the oleate most in demand. I find it a most satisfactory preparation. When the proper amount of oxide has been added, and the B.P. directions followed, it often takes seven days—sometimes longer—to form a solution, which, in my opinion, indicates that it is not a true oleate. The oleic acid which is in excess soon begins to take up the oxygen, reducing the mercury first to suboxide, which is shown by the change of colour from light brown to slate colour, and immediately the whole of the oleate becomes metallic mercury. It takes long to make, and after it is made it is not what it is represented to be. This instability causes trouble to the dispenser, as it prevents him sending out the article always of the same consistency, appearance, and therapeutic strength.

The only way out of the difficulty that I can see would be a modification of the B.P. process, whereby you could make it in a short time. A method that I have found to answer is to place the mortar in a pan of boiling water, placing the weighed oleic acid in the mortar, and maintaining the heat until the oleic acid is about  $100^{\circ}F$ , then dusting the oxide of mercury into it, afterwards stirring until a solution is effected, which takes about fifteen minutes. The oleic acid must not be subjected to a great heat, as heat favours the absorption of oxygen, causing it to smell disagreeably and become rancid. I found that oleate of mercury 10 per cent. made in this manner would keep a month and retain its characteristics. A stronger solution will keep much better, as the excess of oleic acid is the cause of deterioration, owing to its reducing properties.

A true oleate should be made by precipitation and diluted as required. A good method is to make a saturated solution of Castile soap in water, allowing it to stand for a day to get rid of the sodium palmitate, which will deposit, decanting the clear solution and filtering. The filtrate will be composed principally of sodium oleate with some palmitate. Make a solution of a neutral salt of the metal, and precipitate the oleate by reacting on the soap solution. The solution of the metal must not be acid, or we should get some free oleo-palmitic acid. In the case of mercury we should require to keep it boiling briskly for two or three minutes to aggregate the particles. This precipitate—an oleo-palmitate of the metal—can be strained and adherent water got rid of by evaporation in the water-bath.

The oleate may be separated from the palmitate by dissolving out with petroleum ben-