alkyl nitrate:²²

 $RO_2 + NO \rightarrow RONO_2$ (6-39) The peroxyalkyl-NO₂ reaction proceeds principally by $RO_2 + NO_2 \rightarrow RO_2NO_2$ (6-40) The peroxynitrate may thermally decompose according to

 $RO_2NO_2 \rightarrow RO_2 + NO_2$ (6-41) Measured rate constants for the RO_2-NO_2 reaction and the RO_2NO_2 decomposition are not currently available.

Peroxyacyl nitrates have been recognized as important components of photochemical air pollution.²⁷ Peroxyacetyl nitrate (PAN) exists in equilibrium with the peroxyacyl radical and NO₂:

 $\begin{array}{cccc} & & & & & \\ & & & & \\ & & & \\ & & & \\$

The acetyl radical will rapidly decompose as follows: (6-43)

 $CH_3CO \rightarrow CH_3 + CO_2$ (6-44)

followed by:

$CH_3 + O_2 \rightarrow CH_3O_2$	(6-21)
$CH_{3}O_2 + NO \rightarrow CH_{3}O + NO_2$	(6-22)
$CH_3O + O_2 \rightarrow HCHO + HO_2$	(6-23)
$HO_2 + NO \rightarrow OH + NO_2$	(6-24)



I-65