

	Calculated for RbSbCl_4 .		Found.
Rb	85.2	24.60	24.23
Sb	119.6	34.54	35.05 and 35.07
4Cl	141.48	40.86	40.62
	<hr/> 346.28	<hr/> 100.00	<hr/> 99.90

The formula of this salt is therefore RbSbCl_4 .

Summary.

The following is a list of the compounds of antimony chloride, with the chlorides of rubidium and caesium obtained in this investigation. The formula of the first of the rubidium salts must be considered as somewhat doubtful.



These formulas show that the elements in question have a marked tendency towards the formation of complex double chlorides. The most important conclusion, however, to be drawn from the present investigation is that neither the salt described by Godeffroy as $\text{SbCl}_4 \cdot 6\text{CsCl}$ nor that described as $\text{SbCl}_4 \cdot 6\text{RbCl}$ corresponds in composition to the formula proposed by him.

THEORETICAL.

There are now two well-established exceptions to the general law in regard to the composition of the double chlorides, namely, the salts $\text{CuCl}_2 \cdot 2\text{KCl}$ ¹ and $\text{CdCl}_2 \cdot 4\text{KCl}$ ². It is to be noticed, however, that the structure of the salt $\text{Rb}_3\text{Sb}_2\text{Cl}_{10}$ cannot be represented on the same general system as that of the ordinary double chlorides, where two chlorine atoms are supposed to be analogous in function to one oxygen atom in the oxygen salts. This statement holds true even if the formula of this chloride be somewhat simpler than that here proposed. Hence this compound is perhaps quite as exceptional as the two simpler salts mentioned. In view of these facts the author offers the following formulas as suggestions of the possible structure of these compounds. Some

¹ Mitscherlich, Ann. chim. phys. [3] 73, 384.

² C. v. Hauer, Wien. Akad. Ber. 15, 23. These results have recently been confirmed by Dr. G. M. Richardson.