replaceable by a metal, and is the type of monobasic vinic acids, while water is the type of bibasic acids. (*Laurent, Recherches sur les Combinaisons azotées. Ann. de Chim. et de Phys., Nov.* 1856.)

In a review of that remarkable essay, published in this Journal for Sept., 1848 (vol. vi, p. 173), I suggested that this view was "susceptible of still farther extension, and that we may include in the same type all those saline combinations (acids) *which contain oxygen.*" I referred to the hypochlorites Cl O, M O, as derivatives of the type H_2O_2, in which Cl replaces H, being $(ClH)O_2$, and $(ClM)O_2$, while anhydrous hypochlorous acid is Cl_2O_2, the result of a complete substitution. "In the same manner nitric acid, NHO_3, is a monobasic salt (i. e. acid), corresponding to water in which NO_2 is substituted for H, as in many organic compounds; we have then $(NO_2, H)O$ and $(NO_2, M)O$;" or $(NO_4, H)O_2$ in the notation adopted above. "As an adaptation of this idea to bibasic compounds, sulphuric acid, SH_2O_4, is to be regarded as water in which SHO_3 replaces H; thus $(SHO_3, H)O$. As the replacing elements contain an equivalent of hydrogen which is saline (i. e. replaceable by a metal), the acid is bibasic. When the hydrogen in SHO_3 is replaced by a metal, we have a class of acid sulphates like $(SKO_3, H)O$. The complete replacement of hydrogen in the original type yields $(SHO_3)_2O$, which is the Nordhausen acid commonly represented by $2SO_3, H_2O$. This latter compound as Gerhardt has shown, corresponds to the anhydrous bisulphate of potash."

"The tribasic acids may equally be reduced to the same type, if we conceive the elements which replace one equivalent of hydrogen, to be bibasic instead of neutral or monobasic; phosphoric acid, PH_3O_4 is $(PH_2O_3, H)O$."

"The primitive saline type is then essentially bibasic, and is presented in its most elemental form in water, while the simplest type of the monobasic salt, which is a derivative of