operating on teeth of poor structure (which are usually extremely sensitive) it renders this ordeal less painful. Internally it will always prove of great value as a systemic antacid in conjunction with its local use when oral hyperacidity is due to continued eating of acid fruits or the habitual taking of wines at meals, exaggerating a pre-existing tendency in this direction or tending to prevent the normal slightly alkaline reaction of the saliva. (1 will say to you professionally that in that form of dyspepsia characterized by a sour taste in the mouth, red tongue, acid eructations and pyrosis, and aggravated by foods containing too much fat, as well as ordinary cases of gout, rheumatism and gravel arising from an excess of lactic, lithic or uric acid, a dessert to a tablespoonful taken in an equal quantity of water three or four times a day will be attended with the most immediate and satisfactory results.)

In treatments where you are called upon to use acids of different kinds about the mouth, and desire to limit or counteract their action, it is prompt and most agreeable, notably with aromatic sulphuric acid subsequent to the treatment of pyorrhœa alveolaris by this agent. Many persons, dentists as well as others, are in the habit of taking a glass of milk, naturally or artificially prepared, and here its local employment is invaluable in preventing lactic fermentation induced by the heat of the mouth, which, if continued, frequently results in excessively sensitive teeth. The importance of the care of the teeth during pregnancy cannot be overlooked or That superacidity so provoking and dislightly considered. tressing to the patient and so frequently the prime cause of decay when the teeth are kept constantly bathed with an acid saliva, and frequently flooded with acid regurgitations from the stomach, locally and internally employed in these cases the magnesium hydrate has proven promptly effective and agreeable.

It has been asked, "What is the film which is seen to coat the teeth after an application of magnesium hydrate?" The answer is very simple inasmuch as the magnesium exists in the form of a hydrate and is a fluid; so being in the best possible form both physically and chemically for forming new compounds, it is considered that after the excess of acid present has been neutralized the remaining portion of the magnesium hydrate combines in a somewhat loose chemical way with the mucoid or albuminoid constituents of the saliva (which as a rule already coats the teeth), giving us an indefinite compound which may be called, for convenience, magnesium albuminate. Now, this being loosely combined, as soon as more acid is poured out from the glands enough of the alkali is liberated from the albuminous compound by reason of its greater affinity for the acid (which is usually acid sodium phosphate) to neutralize the acid present, such process going on until all the magnesium present is used, which requires several hours as above mentioned.