TABLE I.

Grams Ag employed.	$N/100$-$AgNO_3$ used to titrate back, in c.c.	Nett $N/10$-NaCl used, in c.c.	Equiv. to grams Ag.	Diff. in Ag found, in grams.	Error per cent.
1·0	72·00	92·8926	1·0032	0·0032	+0·32
1·0	76·20	92·4726	0·9987	0·0013	−0·13
1·0	75·05	92·5876	0·99994	0·00006	−0·006
1·0	73·20	92·7726	1·0019	0·0019	+0·19
1·0	72·30	92·8626	1·0029	0·0029	+0·29

Average error = +0·266 per cent. and 0·068 per cent. Mean error = 0·198 per cent.
Average time taken = 27 minutes.

To test the usefulness of this apparatus as applied to the practical estimation of silver in bullion, the usual conditions under which this is done by the Gay-Lussac method were adhered to, namely, 1 gram of pure silver dissolved in about 15 c.c. of nitric acid, 100 c.c. of $N/10$ sodium chloride added, and the excess titrated with $N/100$ silver nitrate. (When necessary, any excess of silver solution added was titrated with $N/100$ sodium chloride.)

The following table contains the results of these experiments, and the time taken:

TABLE II.

Grams Ag employed.	$N/100$-$AgNO_3$ used to titrate back, in c.c.	Nett $N/10$-NaCl used, in c.c.	Equiv. to grams Ag.	Diff. in Ag found, in grams.	Error per cent.
1·004	72·00	92·8926	1·0032	0·0008	−0·08
1·000	73·30	92·7626	1·0018	0·0018	+0·18
0·9994	74·75	92·6176	1·00027	0·00087	+0·087
1·0018	73·80	92·7126	1·00129	0·0005	−0·05
1·0034	73·00	92·7926	1·0021	0·0013	−0·13
1·0030	72·30	92·8626	1·0029	0·0001	−0·01
1·0035	72·30	92·8626	1·0029	0·0006	−0·06

Average error = −0·066 per cent. and +0·133 per cent.
Mean error = −0·067 per cent. Average time taken = 30 minutes.

It may be added that no excessive shaking up of the mixture is required other than sufficient to ensure proper mixing after each addition of the reagent.

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