

F.J. Toole Memorial lectures to be held

Ronald J. Gillespie of McMaster University will give the seventh series of F. J. Toole Memorial Lectures at UNB, Oct. 6, 7, and 8.

Born in England and educated at University College, London, Dr. Gillespie is known for his early work with the late Ronald S.

Nyholm in developing the Valence Shell Electron Repulsion Theory of Molecular Shapes.

This theory is so simple

and easily understood that it is studied by first-year chemistry students the world over. It assists the chemist in designing new

molecules with useful properties, by quickly and accurately predicting the shapes synthetic molecules will take.

The three Toole Lectures are scheduled for 7:30 p.m. Monday, Oct. 6, and 4:00 p.m. both Oct. 7 and 8. The first lecture is open to the public, and will be of particular interest to science teachers, according to UNB chemist Jack Passmore.

All lectures will be held in room 303, F.J. Toole Hall, the UNBF chemistry building.



BACKGROUND ON TOOLE LECTURES

The lecture series, like the building, honors the late Frank Toole, distinguished professor of chemistry and UNB's first dean of arts and science, who retired in 1965 from the post of university vice-president (academic).

Dr. Gillespie follows much eminent Toole lecturers as Nobels Laureates, Linus Pauling and Vladimir Prelog.

His current research as McMaster includes studies of strong acids, fluorine chemistry, and the preparation, characterization and understanding of cluster compounds of the main group elements. He has published more than 250 scholarly papers.

Dr. Gillespie's numerous honors include election as a Fellow of the Royal Societies of Canada and London, the 1967 Canadian Centennial Medal, the American Chemical Society Award for Distinguished Service in the Advancement of Inorganic Chemistry, and the Silver Jubilee Medal in 1978.

The title of his Monday evening lecture is 'The Prediction of Molecular Geometries by Valence Shell Electron Pair Repulsion Theory'; on Tuesday his topic is 'Preparation and Structure of the Polyatomic Cations of Group VI Compounds of the Metals and Non-Metals.'

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