

<p>ESTER OF J. WISLICENUS.</p> <p>Hagemann and Callenbach. Ber. d. deutsch. chem. Ges. 30, 639.</p>	<p>(1) $\begin{array}{l} \text{CH}_3\text{C} \begin{array}{l} \diagup \text{CH} \cdot \text{C}(\text{OH}) = \\ \diagdown \text{CH}_2 \cdot \text{CH}_2 \end{array} \text{C} \cdot \text{COOC}_2\text{H}_5 \end{array}$</p> <p>(2) $\begin{array}{l} \text{CH}_3\text{C} \begin{array}{l} \diagup \text{C} \text{H} - \text{CO} \\ \diagdown \text{CH}_2 \cdot \text{CH}_2 \end{array} \text{CH} \cdot \text{COOC}_2\text{H}_5 \end{array}$</p> <p>(1) Fluid, B.P., 150–152° at 22 mm. (2) Fluid, B.P., 148–152° at 22 mm.</p>
<p>DIBENZOYLMETHANE.</p> <p>Wislicenus, Löwenheim & Wells. Ber. der sächs. Akad. (Sitzung v.) 1, March, 1897.</p>	<p>$\begin{array}{l} \text{CH} = \text{C}(\text{Oid}) \cdot \text{C}_6\text{H}_5 \\ \text{CO} \cdot \text{C}_6\text{H}_5 \end{array}$ $\begin{array}{l} \text{CH}_2 \begin{array}{l} \diagup \text{CO} \cdot \text{C}_6\text{H}_5 \\ \diagdown \text{CO} \cdot \text{C}_6\text{H}_5 \end{array} \end{array}$</p> <p>Solid, M.P., 77.5–78°. Solid, M.P., 77.5–78°.</p>
<p>PHENYLNITROMETHANE.</p> <p>Hantzsch & Schultze. Berd. d. deutsch. Ges. 29, 699. 2251. Rec. trav. chim. 14 46, 365.</p>	<p>$\text{C}_6\text{H}_5 \cdot \text{CH} = \text{N} \begin{array}{l} \diagup \text{O} \\ \diagdown \text{OH} \end{array}$ $\text{C}_6\text{H}_5 \cdot \text{CH}_2 \cdot \text{NO}_2$</p> <p>(Iso-Ph.) Fluid. Solid, M.P., 84°.</p>

It will now be necessary to go into the grounds for assuming that the differences in the compounds are due to tautomeric change and not to any other form of isomerism which plays a part in organic chemistry.

The proofs may be divided into two sections; physical and chemical.

Of the two classes, that one which is the more likely to give certain results is the first, for it assumes that in the process of physical examination no change will take place in the compounds under investigation. In the ease of chemical reaction, one can never be certain that substances, especially those which are susceptible to change, will not undergo under the influence of heat, or cold, solvents, foreign substances in proximity, etc., changes which will render the results, if not valueless, at least uncertain.

The physical methods which lend themselves to the constitutive examination of substances are few in number, but they have in this direction been exceedingly useful. They are:

1. The determination of the molecular refraction.
2. The determination of the molecular magnetic rotation.
3. The determination of the molecular volume.
4. The selective absorption.
5. The absorption for electric oscillations.

The first three of these methods are mainly the result of the representative work of Brühl, W. H. Perkin, sr., and of I. Traube. They are in contradistinction to the latter two, quantitative, and admit in many cases of an accurate judgment being made of the constitution of a compound. They rest on the assumption that the molecular properties of organic substances are the sum of their atomic functions. The spectrometric investigations of Brühl, who has determined the molecular refractions of