

pond to the formula proposed by Godeffroy. The substance was prepared under varying conditions, being crystallised out either by slow cooling of the solution, or by sudden cooling, or by spontaneous evaporation. In the first experiment the two chlorides were mixed in the proportions required by the formula of Godeffroy. In the other cases no special care was taken in regard to the proportions. The composition of the salt formed did not vary appreciably. All the analyses made, in this part of the investigation, are here given, except one attempt to determine rubidium as nitrate, which was quite untrustworthy. The variations from the normal composition are no doubt due to experimental errors, or to the presence of slight impurities in some specimens of the salt.

I. 0.5720 gram gave 0.1902 gram Sb_2S_3 (23.73 per cent. Sb), and 0.8538 gram AgCl (36.91 per cent. Cl).

II. 0.4174 gram gave 0.1401 gram Sb_2S_3 (23.96 per cent. Sb), and 0.2297 gram RbCl (38.89 per cent. Rb).

0.3190 gram gave 0.4772 gram AgCl (36.99 per cent. Cl).

III. 0.4028 gram gave 0.1348 gram Sb_2S_3 (23.89 per cent. Sb), and 0.6007 gram AgCl (36.88 per cent. Cl).

0.6881 gram gave 0.3769 gram RbCl (38.71 per cent. Rb).

IV. 0.2728 gram gave 0.4071 gram AgCl (36.90 per cent. Cl).

V. 0.3359 gram gave 0.1125 gram Sb_2S_3 (23.90 per cent. Sb), and 0.1850 gram RbCl (38.92 per cent. Rb).

A tabular view of these results will make the matter clearer.

	Calculated ¹ for $\text{SbCl}_2 \cdot 6\text{RbCl}$.	Found by Godeffroy.	Found by the Author.			
			I.	II.	III.	IV and V.
Sb	12.60	13.10	23.73	23.96	23.89	23.90
6Rb	53.86	53.06	...	38.89	38.71	38.92
9Cl	33.54	33.45	36.91	36.99	36.88	36.90

These figures show that the salt obtained in the present investigation is a definite chemical compound, which does not correspond in composition to the numbers calculated from Godeffroy's formula. That this salt is really identical with that obtained by Godeffroy, in spite of the great difference between the analytical results in the two cases, is proved by several considerations. This salt is obtained with the greatest ease, much more readily than any other

¹ These figures differ slightly from those calculated by Godeffroy, owing to the use of different atomic weights in the two cases.