

read, not as "equal to," (which would suppress the essential idea sought to be conveyed), but as "yields" or "yield" so and so. For examples of chemical formulæ and equations, see foot notes on pages 35, 36, 37.

List of some Chemical Compounds mentioned in this work, with their formulæ.

Water	H_2O
Silica (quartz, sand)	SiO_2
Silicic Acid	$2H_2O, SiO_2$ or H_4SiO_4
Carbon Dioxide (Carbonic Acid Gas)	CO_2
Sulphuric Acid (Oil of Vitriol)	H_2SO_4
Phosphoric Oxide (Anhydride)	P_2O_5
Phosphoric Acid	$3H_2O, P_2O_5$, or H_3PO_4
Calcium Oxide (burnt lime)	CaO
Calcium Hydrate (slacked lime)	..	CaO, H_2O or CaH_2O_2
Potassium Oxide	K_2O
Potassium Hydrate	K_2O, H_2O or KHO
Sodium Oxide	Na_2O
Ammonia	NH_3
Ferrie Oxide	Fe_2O_3
Sodium Chloride (common salt)	$NaCl$
Calcium Carbonate (marble, limestone)		CaO, CO_2 or $CaCO_3$
Potassium Nitrate (Saltpetre)	..	K_2O, N_2O_5 or KNO_3
Calcium Sulphate (Plaster, anhydrous)		CaO, SO_3 or $CaSO_4$
Tri-Calcic Phosphate (bone earth)		$3CaO, P_2O_5$ or $Ca_3P_2O_8$
Bi-Calcic Phosphate (Reduced Phosphate)		$2CaO, H_2O, P_2O_5$ or $Ca_2H_2, 2PO_4$
Mono-Calcic Phosphate (Superphosphate)		$CaO, 2H_2O, P_2O_5$ or $CaH_4, 2PO_4$
Aluminium Silicate, hydrated, (Silicate of Alumina, Clay)	$Al_2O_3, 2SiO_2, 2H_2O$
Aluminium and Potassium Silicate (Double Silicate of Alumina and Potash)	$K_2O, Al_2O_3, 6SiO_2$