

read, not as "equal to," (which would suppress the essential idea sought to be conveyed), but as "*yields*" or "*yield*" so and so. For examples of chemical formulæ and equations, see foot notes on pages 35, 36, 37.

List of some Chemical Compounds mentioned in this work, with their formulæ.

Water	H_2O
Silica (quartz, sand)	SiO_2
Silicic Acid	$2H_2O, SiO_2$ or H_4SiO_4
Carbon Dioxide (Carbonic Acid Gas)	CO_2
Sulphuric Acid (Oil of Vitriol)	H_2SO_4
Phosphoric Oxide (Anhydride)	P_2O_5
Phosphoric Acid	$3H_2O, P_2O_5$, or H_3PO_4
Calcium Oxide (burnt lime)	CaO
Calcium Hydrate (slacked lime) ..	CaO, H_2O or CaH_2O_2
Potassium Oxide	K_2O
Potassium Hydrate	K_2O, H_2O or KHO
Sodium Oxide	Na_2O
Ammonia	NH_3
Ferrie Oxide	Fe_2O_3
Sodium Chloride (common salt)	$NaCl$
Calcium Carbonate (marble, limestone)	CaO, CO_2 or $CaCO_3$
Potassium Nitrate (Saltpetre) ..	K_2O, N_2O_5 or KNO_3
Calcium Sulphate (Plaster, anhydrous)	CaO, SO_3 or $CaSO_4$
Tri-Calcic Phosphate (bone earth)	$3CaO, P_2O_5$ or $Ca_3P_2O_8$
Bi-Calcic Phosphate (Reduced Phosphate)	$2CaO, H_2O, P_2O_5$ or $Ca_2H_2, 2PO_4$
Mono-Calcic Phosphate (Superphosphate)	$CaO, 2H_2O, P_2O_5$ or $CaH_4, 2PO_4$
Aluminium Silicate, hydrated, (Silicate of Alumina, Clay)	$Al_2O_3, 2SiO_2, 2H_2O$
Aluminium and Potassium Silicate (Double Silicate of Alumina and Potash)	$K_2O, Al_2O_3, 6SiO_2$