

trolyte. In these experiments equation IX assumes the form:

$$\frac{\{Cu^{++}\}}{\{Cu^+\}^2} = \frac{\frac{(2Q-P)}{63.6} \left(\frac{1}{V} + \frac{10^3 l}{k_2 A t} \right) + 0.171 \times 10^{-3}}{\left[\frac{2Y(P-Q)}{63.6} \left(\frac{1}{V} + \frac{10^3 l}{k_1 A t} \right) \right]^2} \quad (IX a)$$

In the Introduction the qualitative effect of changes in rate of stirring was discussed. Table XII gives the data for three electrolyses which were carried out with lower rates of stirring than that used in the other experiments. Under R is given the rate of stirring in revolutions per minute, and in the last

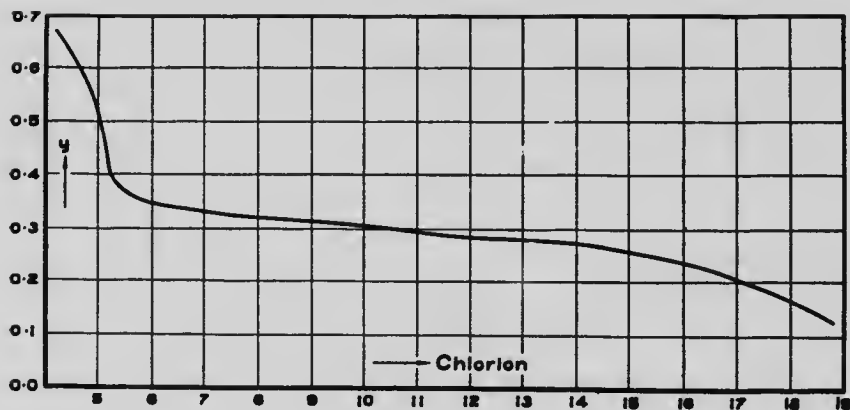


Fig. 2

column of the table, under C, is given the number of experiment in Table X in which all the other variables the same and which may thus be used for comparison. The experiments show that in accordance with the requirements of the theory the proportion of cuprous salt formed is less when the rate of stirring is decreased.