very slightly, due no doubt to the oxidation of the bromion or to the reduction of the chlorate.

Noyes and Scott worked with the potassium salts of bromic acid and hydriodic acid in presence of a constant amount of hydrochloric acid; their constant agrees very closely with my own.

Magnanini's Measurements

TABLE XIV

$HBrO_3$, 1.85; HI, 11.11

t	x	$K \times 10^{13}$
5	0.82	142
12	1.44	119
13	1.46	116
21	1.84	100
25	2.05	99
41	2.72	93
52	2.96	91
84	3.69	88
91	3.80	89
175	4.82	86

TABLE XV

$HBrO_3$, 1.85; HI, 11.11; HCl, 11.11

t	x	$K \times 10^{13}$
2	0.82	99
3	1.05	87
9	2.35	84
15	3.27	85
17	3.38	81
28	4.38	84
31	4.50	83
35	4.68	84
40	5.10	86

TABLE XVI

$HBrO_3$, 1.85; HI, 11.11; HCl, 22.22

t	x	$K \times 10^{13}$
2	1.43	85
6	2.98	83
12	4.19	80
17	4.93	80
18	5.06	80
20	5.28	81

TABLE XVII

$HBrO_3$, 1.85; HI, 11.11; HCl, 33.33

t	x	$K \times 10^{13}$
2	2.03	83
4	3.17	82
6	3.93	81
8	4.52	81
10	4.90	80
11	5.08	79

TABLE XVIII

$HBrO_3$, 1.85; HI, 11.11; HCl, 44.44

t	x	$K \times 10^{13}$
2	2.55	68
4	3.94	70
6	4.72	70
7	5.15	72
8	5.33	72

TABLE XIX

$HBrO_3$, 1.85; HI, 11.11; HNO_3, 22.22

t	x	$K \times 10^{13}$
2	1.37	82
6	2.88	78
12	4.19	77
17	4.84	77
19	5.14	78