

Radford, as above noticed, I have strong sympathy, feeling convinced that the longer the placenta is allowed to remain, the greater will be the danger to the patient from hæmorrhage and nervous exhaustion, and proportionately difficult will be found the reductive process. We must not altogether despair, however, if the time allowed to elapse after the accident be great before we succeed in replacing the organ, as Dr. Melrose reports a case reduced eighty hours after the inversion occurred.

Finally, I would say that in every case of inversion, of whatever degree, an attempt should be made to remove the placenta before we try to restore the organ, for the following reasons :

1st. If the placenta be returned adherent, the difficulty of its detachment afterwards, as well as the dangers of nervous shock to the mother, will be much greater than before.

2nd. The obstructions encountered in an attempt to replace the uterus, even without the placenta, are nearly always great, and will be incomparably greater if it be allowed to remain, with the utero-peritoneal cavity rapidly filling with appendages, omentum, intestines, and blood.

3rd. With the placenta promptly removed, there is *much less* fear of hæmorrhage, both before and after reduction, as in the case I have particularly alluded to, than if it be allowed to remain, and the entire mass returned together.

THE SULPHO-CARBOLATES.

By F. H. WRIGHT.

Within the last few months, the attention of the medical profession has been directed to the beneficial results, claimed to have been obtained from the use of the sulpho-carbolates, particularly in zymotic diseases. Their superiority over carbolic acid is said to have been satisfactorily proved, inasmuch as they possess all the *advantages* attending the use of this acid, without any of the difficulties or dangers arising from its administration. The salt, also, does not possess the nauseous odour and taste so much complained of in carbolic acid—even when largely diluted—and is consequently much more easily administered, particularly to children.

Although this double acid has but very re-

cently attracted much attention, it is by no means a new preparation; sulpho-phenisic acid, and some of its salts, being first mentioned by Laurent, in the October number of the *Annales de Chimie et de Physique* for 1841.

Sulpho-carbolic acid is readily prepared by mixing two parts by volume of carbolic acid with one of sulphuric, and heating in a suitable vessel for five minutes at a temperature of 290° F.

The sulpho-carbolates are obtained tolerably pure by diluting the above mixture, when cool, with six or eight volumes of water; and then neutralizing with a base. However, I have only once been able to obtain them absolutely pure in this way; barium chloride always causing a turbidity in the solution, thereby showing the presence of either free sulphuric acid, or the sulphate of the base employed.

The use of the double volume of carbolic acid may be readily understood by referring to the atomic weight of the acids employed :

Carbolic Acid.	Sulphuric Acid.
C ₆ = 72	H ₂ = 2
H ₈ = 6	S = 32
O = 16	O ₄ = 64
94.	98

Here we find that the combining equivalents are nearly equal; but sulphuric acid is nearly twice as heavy as an equal volume of carbolic—hence, the use of the double volume of the latter is very nearly correct.

The so-called carbolic acid, like other alcohols containing a large amount of carbon, seems to form the sulpho acid with remarkable facility, nearly the whole of the sulphuric acid being converted into sulpho-carbolic. The process is therefore, much simpler than that for preparing sulphovinates and sulphomethylates; where it is necessary to form the barium salt first, and then decompose this with sodium or other soluble carbonate or sulphate. To obtain the sulpho-carbolates pure, they should be prepared in this way, as by the process above described, they usually contain a trace of sulphate.

Impurities.—The impurities usually found in these salts, are the following:—(1) The sulphate or carbonate of the base employed. (2) Free carbolic acid. (3) Free sulphuric acid. The last rarely occurs.