Drug Review.

Business during the past month has been fairly active for this season of the year, and prospects for a good fall trade are good.

There have been no great changes in values.

Carbolic Acid in particular and disinfectants in general have advanced in price on account of large demands in countries where cholera exists, or where it may be expected.

Silver Nitrate is lower than it has been for years.

Liquor Ammonia has been reduced in price.

Oil Pennyroyal is higher.

Canary Seed has advanced.

Alcohol has advanced 5 cents a gallon in small lots.

Tannic Acid is easier in price. Tartaric Acid is very low. Cocaine and Sulfonal are easier.

Make your prices as reasonable as your expenses and quality of material will permit.

Tur use of antipyrin is said to cause blackening of the teeth in some individuals, especially when the enamel is imperfect. The discoloration may be removed with dilute hydrochloric acid.

Reaction Between Powdered Borax, Glycerin and Sodium Bicarbonate.

J. U. LLOYD.

Some months ago a druggist handed me a prescription containing ingredients about as follows: Powdered Borax, 1 ounce; bicarbonate of sodium, 1 ounce, carbolic acid 10 drops; water and glycerin, of each, 2 ounces, asking me what would be the result of the compounding of the prescription. It did not occur to me that any unusual reaction would take place, and I was surprised on being informed that the mixture had been compounded and had exploded in the bottle. Upon experimenting with the ingredients I found that the water and carbone acid were passive, and that the prescription could be filled without the glycerin without visible reaction. The addition of glycerin, however, produced violent effervescence by reason of the liberation of carbonic acid gas, and it was found that a mixture of bicarbonate of sodium and powdered borax reacted upon each other in the presence of glycerin, producing sodium borate and carbon dioxide, a fact that had previously escaped my observa-The matter was mentioned to Prof. Norton, then President of the Cincinnati Chemical Society, and he agreed to look up the literature on the subject (if any existed), the fact that such reac-

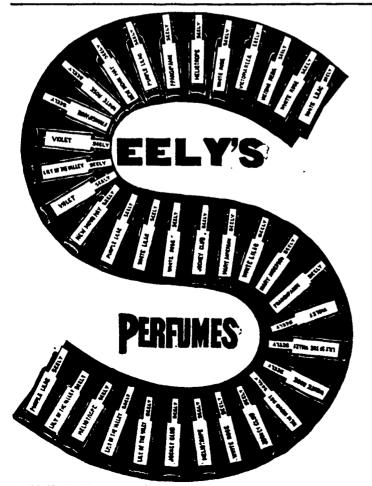
tion would result from these substances having also escaped his attention. In compliance he mailed me the following abstract from the Boston Journal of Chemistry, December, 1877, from which it seems that the combination had been studied previously. The phenomenon may be of interest to others in pharmacy, for I have reason to believe that the reaction has been overlooked by most of

those who fill prescriptions.

Mr. M. W. 1bes, of the Hopkins University, gives the following as his explanation of the effervescence on mixing glycerin, borax, and sodium bicarbonate.

"Since glycerin is one of the best solvents known, and also since glycerin dissolves more carbonate of soda than of any other salt, therefore when these salts come into solution together there will be a displacement of one molecule of carbonic acid by one molecule of boracic acid, and the resulting product will be two molecules of normal or neutral borate of soda, because when boracic acid is in solution it is a stronger acid than carbonic acid (see Gmelin's 'Handbook of Chemistry'). Furthermore, the readiness with which the chemical action takes place is partly due to the fact that boracic acid neutralizes the alkalies imperfectly, a fact clearly substantiated by borates having an alkaline reaction."

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