

Bromine water and then ammonia: green, flocky, precipitate, soluble in excess of ammonia to form emerald green solution; neutralized with dilute hydrochloric acid, the solution becomes blue, and with excess of acid, violet or red.

Potassium ferrocyanide: dark red coloration.

STRYCHNINE, $C_{21}H_{22}N_2O_2$.—White glistening rhombic prisms, very difficultly soluble in water and in alcohol, easily soluble in dilute acids.

Sulphuric acid conc.: dissolves strychnine to form a clear solution. If to a small portion of this solution a fragment of potassium dichromate be added, a violet coloration, surrounding the dichromate, may be observed. This coloration is produced also (but less distinctly) by potassium permanganate, and by red lead.

Potassium ferricyanide: yellowish green coloration.

THEBAYNE, $C_{19}H_{23}NO_3$.—White quadratic scales, insol. in water, easily in alcohol and in acids.

Sulphuric acid conc.: solution with blood red color, gradually becoming yellowish red.

Ferric chloride: no color reaction (distinction from morphine).

Nitric acid, sp. gr. 1.4: solution with yellow color.

VERATRINE, $C_{32}H_{40}NO_9$.—Colorless prisms; insol. in water, soluble in alcohol and in acids.

Sulphuric acid conc.: yellow coloration, becoming bright red on standing. Upon addition of bromine water this solution becomes purple.

Hydrochloric acid conc.: colorless solution cold, colored deep red upon boiling.

CAFFEINE OR THEINE, $C_8H_{10}N_4O_2H_2O$.—White, glistening, silky needles; insol. in alcohol, soluble in water and in acids.