

thenar space; the middle palmar space; the hypothenar space; the radial bursa; the ulnar bursa. Through the last two spaces infections usually spread from the hand to the forearm.

The line of treatment lies in early adequate and properly placed incisions; drainage, if any, by rubber tissue; and the use of Bier's hyperemia. *Ther. Gazette.*

HAZARDS IN HANDLING GASOLINE

During 1913 in the United States 1,040 persons were burned to death and 3,120 persons injured (a weekly average of twenty persons killed and sixty persons injured) on account of gasoline fires. In harmony with experience respecting human ways and human indifference, the increasing use of gasoline and other liquid fuels which are dangerous because of their ready inflammability will probably increase the number of accidents which cause loss of life as well as great damage to property. Miscellaneous hazards in the use of the inflammable products exist which lead to leakage of these volatile substances and their ignition by spontaneous and unsuspected methods, as well as those due to obvious carelessness. A feature, however, which is not readily comprehended and which is a serious source of danger is the fact that gasoline or gasoline vapor burns, but a mixture of the vapor and air is highly explosive. The Bureau of Mines illustrates the situation in these words: If one takes the cover off a full pail of tightly enclosed gasoline and applies a match to the surface, the gasoline will flare up and burn as long as the gasoline lasts. On the other hand, if one puts a few drops of gasoline in a small tightly enclosed pail, waits a few minutes, and then introduces a flame or an electric spark, a violent explosion will most likely result. In the first case the vapor burns as fast as it comes from the gasoline, and mixes with the oxygen of the air; in the second, the oil vaporizes in the pail and mixes uniformly with the air in it to form an explosive mixture, and on ignition explodes. Consequently, when one hears of a disastrous gasoline explosion one may be sure that it resulted from the mixing of the vapor from the gasoline with air in the proportions necessary to form an explosive mixture. The behaviour of illuminating gas, which burns quietly when liberated alone, but explodes when a mixture with air is heated, is quite analogous. The public has been slow to appreciate these distinctions, and hence they deserve emphasis. At ordinary temperatures, air will hold from 5 to 28 per cent. of gasoline vapor. As gasoline vapor is about three