

the author has convinced himself. But business considerations prevent the further elucidation of this subject."

And further :

"It must be evident to every one familiar with the facts, that the commercial electrolysis of copper on the one hand, and an electrolytic deposition of nickel in our nickel-plating establishment on the other hand, point out a path to follow that is too plain to be neglected.

"And as our chemists find no difficulty in precipitating with the electric current chemically pure copper from a solution containing both copper and nickel, and then, by slightly altering the conditions, precipitating all the nickel in absolute purity from the same solution and with the same current, it would seem that our refiners might reasonably expect to effect the same results on a commercial scale, especially as there is practically no loss of acid in the operation.

"Nor can I see any reason why nearly all our metallic nickel should be offered to the trade in little cubes less than an inch square. Of course this peculiar form has resulted from the practice of the nickel refiners to reduce the oxide of nickel obtained by the methods now in use to metallic nickel. Being mixed with rye meal, as a reducing agent, it is formed into these little cubes, and a number of these packed in crucibles are exposed to a sufficient heat to reduce the nickel to a metal without fusing it. This makes a small porous fragment of metal suitable for solution in acids, and where nickel is to be