the first product of oxidation, preceding anisaldehyde. Anisoine $C^{20}H^{12}O^2$ + 6 HO+4 O=$C^{20}H^{18}O^{12}$. He has examined several of the salts, the silver compound is $C^{20}H^{17}O^{11}$, Ag O.—*Ch. Gaz.* 325.

Anilotic Acid.—Major has examined the acid obtained by Piria by the action of nitric acid upon salicine, to which he gave the above name, and states it to be identical with nitrosalicylic. Piria denies this, and recommends for its formation the following process: Into a stoppered bottle, 1 part of powdered salicine and 6 to 8 parts of nitric acid of 20° B are put, the bottle is closed and placed in a cool place, hyponitrous acid is formed, the fluid becomes green, and after some time crystals of anilotic acid separate. If the process be conducted in an open vessel, the liquid becomes yellow, and helicine is formed. The properties of the acid are described.—*Ch. Gaz.*, 325.

Arachic Acid.—Scheven and Gössmann have described the salts, ether, amide and glyceride of the above acid. Its formula is $C^{40}H^{40}O^4$; the acid is obtained from ground nut oil.—*Ch. Gaz.*, 326.

Ethylamine.—Emil Meyer has described various salts and double salts of this base, with phosphoric, sulphuric and molybdic acids, &c., &c.—*Ch. Gaz.* 327.

Acids in the Animal Organism.—Bertagnini finds that camphoric acid passes unchanged into the urine, the anhydrous acid becomes hydrated, anisic acid passes unchanged, salicylic acid rapidly passes into the urine as indicated by the iron test, but a portion becomes converted into a new compound which he calls salicyluric acid, having taken up the elements of glycocine and lost two equivalents of water. The acid can be separated by evaporating the urine, separating from the salts, acidulating with hydrochloric acid, shaking with ether, evaporating, and recrystallizing. The salicylic acid is removed by heating to 284—302 F in a current of air,—the residue is decolorized and crystallized. The formula is $C^{18}H^9NO^8$.—*Ch. Gaz.*, 325.

Saponification.—Pelouze finds that fats can be readily saponified by the anhydrous oxides or their hydrates in a solid form, if the mixture be heated to 482° F. With suet the soap formed yields from 95 to 96 per cent. of the suet operated on. During the reaction a white smoke is evolved with an odour of burnt sugar, that of acetone is also perceptible. 10 parts of anhydrous lime are sufficient for 100 parts of suet, with 12 or 14 the reaction takes place with much greater facility; but on operating with large quantities it is difficult to keep within bounds so as to prevent decomposition.

Slaked lime in the proportion of 10 to 12 per cent. rapidly saponifies fats at a temperature between 410° 447° F. Two pounds of suet with 120 grammes of slaked lime were saponified in one hour; if the temperature be raised rapidly to 482°, the process may be completed in a few minutes.

This fact promises to be of very great importance to the manufacturers of the so-called stearine candles.

H.