phuric, nitric or acetic acids; but insoluble in hydrochloric acid or alkalies. Its uses are the same as those of the alkaloid. Tannic acid is supposed to be an antidote for strychnia. It will act by combining with the strychnia in solution, and thus render it insoluble; but this is not a sufficient power, as we have seen that the compounds of strychnia are not rendered less powerful by insolubility, as has been considered.

CHROMATE OF STRYCHNIA.—$\overset{+}{St}$ Cr O^3 or (C^{42} H^{22} N^2 O^4) Cr O^3. This is prepared by adding a solution of bichromate of potash to one of a soluble salt of strychnia. The chromate falls down as an insoluble yellowish powder, which appears like broken down cubical crystals. It is insoluble in water, acetic acid, or alcohol; but soluble in sulphuric, nitric, and hydrochloric acids and aqua ammoniæ. Its uses will be similar to those of the alkaloid.

CARBAZOTATE OF STRYCHNIA.—$\overset{+}{St}$ C^{15} O^{15} N^3 or (C^{42} H^{22} N^2 O^4) (C^{15} O^{15} N^3.) This is prepared by dissolving Strychnia in Carbazotic Acid. It crystallizes in small, round, needle-like crystals, which are whitish and aggregate in stars. It is very soluble in water, alcohol, sulphuric, nitric and carbazotic acids; but insoluble in ether or ammonia. Its uses will be similar to those of the alkaloid itself.

BENZOATE OF STRYCHNIA.—$\overset{+}{St}$ (C^{14} H^5 O^2) or (C^{42} H^{22} N^2 O^4) (C^{14} H^5 O^3.) This salt is prepared by dissolving the alkaloid in Benzoic Acid. It crystallizes from the aqueous solution in groups of crystals, having a cubical form; many are pointed at the extremity. It is not very soluble in cold water. It has no importance except in being a salt of strychnia, and its uses will likely be similar to those of the base.

GALLATE OF STRYCHNIA.—$\overset{+}{St}$ *Gal* or (C^{42} H^{22} N^2 O^4) (C^1 H^3 O^5.) This is a soluble salt, and is merely worth mentioning, because this acid does not form an insoluble compound with the base like Tannic Acid, and would serve as a means of distinguishing these two acids. Gallic acid might be a negative proof of Strychnia from this property, which few other elements possess. When the aqueous solution is evaporated, the salt is deposited in colourless scales: these appear in the field of the microscope like thin broken glass fragments. I could not distinguish any distinct crystalline form. It is more soluble than the Benzoate, and is prepared by dissolving the alkaloid in Gallic Acid and evaporating until it takes its peculiar form. It is not very soluble in cold water. One thing a little remarkable about the salt is, that both Strychnia and Gallic Acid are very insoluble in water taken separately; but when