

A review of the various formulæ that have been published in the London and British Pharmacopœias will serve not only to show the relative amounts of ferric oxide directed to be added to the citric acid, but also to indicate (when the examination of commercial specimens is brought forward) that obsolete processes are followed by manufacturers.

As this iron salt is almost universally used it will be interesting at the same time to note the formulæ of the French Codex and the United States' and German Pharmacopœias.

London Pharmacopœia, 1851. (First officinal.)

Sulphate of Iron	12 ozs. Troy.
Carbonate of Soda	12½ ozs.
Citric Acid	6 ozs.
Solution Ammonia ... ('960).....	9 fl. ozs.

British Pharmacopœia, 1864.

Solution of Persulphate of Iron (1·441)	8 fl. ozs.
Citric Acid	5 ozs. (Avoir.)
Sol. Ammonia ... ('960).....	14 fl. ozs.

British Pharmacopœia, 1867.

Solution of Persulphate of Iron (1·441)	8 fl. ozs.
Citric Acid	4 ozs. (Avoir.)
Sol. Ammonia ... ('959).....	19½ fl. ozs.

United States' Pharmacopœia, 1873.

Citric Acid	5 oz. 360 grs. (Troy).
Sol. Persulphate of Iron (1·320)	16 fl. ozs.
Sol. Ammonia ... ('960).....	20 fl. ozs.

French Codex, 1856.

Acid Citric	100 parts.
Hydrated Peroxide of Iron	q. s.
Sol. Ammonia	18 parts.

Add such a quantity of hydrated ferric oxide as will correspond to '53 parts of anhydrous oxide iron.

Pharmacopœia Germanica, 1872.

Citric Acid	2 parts.
Oxide of Iron	q. s.
Then add Citric Acid	1 part.
Sol. Ammonia	q. s. to saturation.

The following table, deduced from these formulæ, will show at a glance the relative amounts of anhydrous ferric oxide to the same amount of citric acid :