<i>Neutions of Polassium Bromale, Indiae, Etc.</i>	um Bromate, Iodide, Etc.	į .	f Potassium	of	Reactions
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$K1: J_{3}, 20.5;$	KBr, 2000;	1, 98.67;	HCl, 95.15	
t	x	x' (mixture)	x' (Sum I and V)	
20	1.28	1.31	1.52	
30	1.87	1.93	2.39	
45	2.75	2.87	3.67	
60	3.40	3.58	4.84	
90	5.31	5.80	7.82	
			/	

TABI VI

[Note: As the concentrations of the reagents have the same effect on the rate of oxidation of bromide¹ as on that of iodide, the values of x' in Tables V and VII were calculated from equation (2).

In calculating x' (mixture) for Tables VI and VIII the concentration of the bromide has been regarded as constant, and x' has been computed by means of a modified form of equation (2), viz.:

$$x' = x + (1/2A + 1/C)x^2 + (x-y)^2/2B$$
 (2 bis)

where y is the value of x in Table V, and x - y is substitut for x in the term $x^2/2B$, y being regarded as the iodine liberat (in experiments VI) by the bromine set free by the action of bromate on bromide, and x - y that liberated by direct action of bromate on iodide.]

TADE VII

 KBrO ₃ , 20.5;	KBr, 2000;	HCl, 190.3	
t	x	· x'	
20	2.40	2.46	
30	3.76	3.91	
45	5.39	5.70	
бо	7.00	7.50	
90	9.71	10.71	

' Judson and Walker: loc. cit.