

Some of the Chemistry class no doubt consider such names as *permanganate of potassium*, *ferrocyanide of potassium*, etc., as difficulties in the pursuit of chemical knowledge. What would they say to studying the properties of the following six chemicals?—

Carboxethylorthoamidoparatoluyamide, nitrometameihylorthouramidobenzoyl, amidometamethylorthouramidobenzoyl, dinitrometamethylorthouramidobenzoyl, diamidometamethylorthouramidobenzoyl, and the diacetyl derivative of diamidometamethylorthouramidobenzoyl.

Lecture 6 was on Oxygen. Oxygen was generated on a large scale and stored for after experiment, some in a large bag, and some in the metal cistern of the pneumatic trough. Priestley's original experiment with oxide of mercury was repeated and explained. To illustrate the characteristic properties of Oxygen in the pure state, many jars of the gas were examined. The principle of the Lime-light was explained, and the Laboratory was flooded with its bright rays. By means of an electric current, water was decomposed into Hydrogen and Oxygen. The graduated test tubes which received the separated gases showed clearly the volume-composition of water:—two vols. of Hydrogen to one of Oxygen. To show how strong is the affinity of these gases for each other, we filled a common Florence oil-flask with the gases in the above proportions. An electric current was then passed through a bit of thin platinum wire contained in the flask. The loud explosion which followed was surprising, when the small volume of the mixed gas is remembered. Oxidation was explained by experiments on the combustion (oxidation) of Magnesium, Carbon, etc.

A noted toper, who was once present at some experiments on water,—similar to ours, was so impressed with them, that he said—"well, I never knew that water was so dangerous, I'll never touch another drop."

In consequence of some enquiries addressed to us from an outport, we made enquiry of the Government as to the