

ing a clear pink or red solution which is strongly alkaline in reaction. To this is added dilute acetic acid¹ (25 per cent) drop by drop; the solution effervesces and becomes deep red in color. The acid is added carefully until the red color entirely disappears, leaving a clear slightly yellow transparent solution which is neutral or even slightly acid to litmus. It is then immediately taken up in a sterile syringe and injected subcutaneously.

By this method a gram of either salt may be obtained in solution in as small a quantity as 1 cc. of fluid. In most of our experiments we prepared the solutions so that 2 gm. would be contained in 5 cc.

This solution must be injected immediately, for if allowed to stand it becomes gradually and progressively alkaline in reaction, so that within one hour it has again taken on a decided red color and a white precipitate has formed. It cannot be heated, for the colorless neutral solution is decomposed even by gentle warming, resulting in an intensely red solution which is strongly alkaline in reaction.

When a dilute solution is desired, water must be added prior to neutralization, or else a weaker solution of acetic acid employed. The addition of even a drop of water to the neutral solution causes a white precipitate to be thrown out, which is re-dissolved upon shaking, but if larger amount are added the precipitate is thrown out and is only dissolved in a great excess of water.

Experiments were made to see if these phenolphthalates possess purgative properties. The neutral solution above described was injected into dogs which were kept in separate metabolism cages and whose diet and water had been carefully controlled and the condition of whose stools was known. Only such dogs whose stools were well formed, were hard, dry and friable, were used. The skin was shaved in spots and the drug injected under aseptic conditions and with antiseptic precautions.

The data relating to these experiments are given in Table I.

¹ Acetic acid must be used in this connection as the addition of any of the mineral acids gives a copious white precipitate immediately upon the neutral point being reached.