on sodium carbonate[1] the hydrated salt is formed; and if a little alcohol be dissolved in the water, it seems to make no difference. But if strong alcohol be used, i.e. a solution of water and alcohol containing only a few percent of water, no hydrate is formed; indeed if a clear crystal of the latter substance be covered with the alcohol it becomes opaque and falls to a powder of sodium carbonate. In this instance the "chemically" combined water resembles pure water in combining with alcohol, and the "physically" combined water (in the "strong alcohol") differs from pure water in not combining with sodium carbonate.

It is natural to ask why there should be a general tendency to exaggerate the difference between these two classes of compounds. An answer is given in the opening paragraphs of the chapter on Reactions.

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1 Prepared by heating washing soda in an evaporating dish over a Bunsen burner.