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NOTE ON "SCRAPS," IN PHARMACEUTICAL JOURNAL,
(p. 383).

BY H. H. CROFT.

The fact that iodine is separated when ferric chloride is added to potassium iodide, has been often noticed and forms a rather disagreeable reaction in the usual method of detecting acids by silver nitrate. Having reduced down the acids to those giving a white precipitate with the silver salt, insoluble in hydric nitrate, the action of ferric chloride is sometimes used for distinguishing or separating into groups hydroferrocyanic and sulphocyanic acids. The yellow or brownish color produced by an iodide is often mistaken by students for the blood-red of the sulpho-cyanate.

That the separation of iodine is owing to a reduction of the ferric salt, may be easily shown by making two solutions of ferric chloride and adding to one a few drops of solution of potassium iodide. On the addition of a little potassium-ferricyanide, the one solution will turn brown, the other dark blue, showing the formation of ferrous salt.

The decomposition may be represented thus :



"Monad's" observation on the precipitate, obtained on adding $\text{Fe}^3 \text{Cl}^6$ to a solution of strychnine in presence of potassium iodide, attracted my attention, and having made a few experiments, I find