boiled together they form a very soluble salt, although they will not unite in cold water. Its uses are most likely similar to those of the alkaloidal base.

There are other salts with the Oxygen acids, which have nothing very remarkable about their composition and which are soluble. They are the Lactate, Citrate, Tartrate, Oxalate, Borate and Phosphate. They are all easily prepared by dissolving the alkaloid in each acid respectively and evaporating the watery solution.

2. Strychnia with the Halogen Radicals.—The halogen elements do not unite with the oxide of a metal or alkali; and, reasoning from the same premises, we might say that each of the alkaloids is the oxide of some base with which we are unacquainted in its free state; and it might be supposed, that, when the haloid radical separates from it, oxygen immediately combines with it, forming the common alkaloid. This theory would do very well with most of the alkaloids; but there would be a difficulty with those which do not contain any oxygen at all, and yet unite with the oxygen and hydrogen acids indiscriminately, as Narcotina, and, according to Artigosa, Conia, although Liebig gives to it one equivalent of oxygen; while Löwig says it does not contain oxygen, and that it is decomposed by Chlorine, Iodine, and Bromine, when brought in contact with them. Upon any other view we must suppose alkaloids are peculiarly constituted bodies, in which the elements of combination unite in opposition to our generally received laws; that the same elements hold inviolate with regard to every other substance, not excepting the alkaloids of the alcohol series, except when they unite with an oxygen acid of these compounds, as the chloric. If, on the other hand, we consider these vegetable bases to be elementary with regard to the halogen, then the oxygen acids must forego their usual law and unite with an unoxygenised base, which they do not do with others. Perhaps the halogen radicals, in combining with these bases, displace an equivalent of hydrogen, as they are known to do in the alcohol series, and also as shown in the formation of the chloride of carbon, thus acting by substitution. However this may be, I am not inclined to think that the chemical laws are at all interfered with in these combinations, although the rationale is unexplained; and, in our present state of knowledge on this point, it would be preferable to describe them as if they were bases of oxygen acids; and, instead of terming them Chlorides, I shall call them Hydrochlorates, and in the same manner as with other elements.

Strychnia and Chlorine — Hydrochlorate of Strychnia.—$\overset{+}{St}$ HCl or (C^{42} H^{22} N^2 O^4) HCl.—This salt is officinal in the Dublin Phar-