

Should such limitations as foster monopoly in the manufacture and sale of such products be removed in the interest of the public good?

5. The relations of pharmacists to public sanitation.

6. Statistics of the present number of pharmacies in proportion to population in various countries, and of imports and exports of crude drugs, medicinal chemicals, and pharmaceutical preparations during the last half-century.

#### SECTION II. — PHARMACEUTICAL EDUCATION AND LEGISLATION.

1. Statistics giving the number of schools or colleges of pharmacy in each country, and the total number of students pursuing pharmaceutical courses.

2. How do the education and the professional and social position of pharmacists compare with those of other professions?

3. What legislation, if any, is at present most needed for the advancement of the best interests of pharmacy?

4. To what extent is official supervision of drug-stores necessary or beneficial?

#### SECTION III. — PHARMACOPŒIAL QUESTIONS.

1. The proper scope of a national pharmacopœia.

2. What improvements, if any, are desirable and practicable in pharmacopœial nomenclature? Is a nearer approach to international uniformity possible?

3. What would be an ideal pharmacopœia?

4. What progress has been made towards the preparation of an international pharmacopœia for potent remedies?

What action, if any, should be taken in reference to this subject?

5. Have the influence and co-operation of pharmacists increased in the work of pharmacopœial revision in the various countries? What proportion of the membership of the pharmacopœial revision committee or commission of your country consists of pharmacists?

6. Should any substance, the manufacture or sale of which is restricted by any patent, copyright or trade-mark, be admitted into any national pharmacopœia? If so, under what conditions?

7. What consideration should determine the introduction into the Pharmacopœia of a new remedy, or the retention or rejection of one already in it?

#### SECTION IV. — GENERAL SECTION.

1. Upon what general plan can a systematic pharmaceutical nomenclature of the complex organic chemicals recently being introduced into the Materia Medica (such as antipyrine, etc.) be constructed?

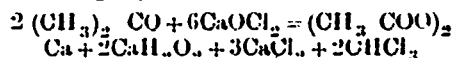
2. In what directions may the pharmacist profitably extend his technical and professional work to render him less dependent upon the purely mercantile part of his business?

Papers upon these and other subjects which may be presented and accepted will be referred to their appropriate sections.

To preserve ammonium carbonate pour a little water of ammonia into the container.

### Ketone Chloroform.

This name was first originated a few years ago to distinguish from chloroform made by the ordinary process, a product prepared independently of ethyl alcohol. The difficulties at first prophesied, with reference to the production of the article in a pure state and at a price which would compete with that of the chloroform made from methylated spirit have been proved to be surmountable; the product is now placed upon the market in a highly pure state, and at a price which has enabled it to become a formidable rival of the ordinary "pure chloroform" of commerce. Of course, it is scarcely necessary to point out that the word "ketone" is a generic name for a large class of compounds characterized by the presence in the molecule of the group  $\text{C}=\text{O}$ , combined with two monivalent alcohol radicals. In all probability, however, the most convenient and available member of the class for the purpose would be methyl-ketone  $(\text{CH}_3)_2\text{C}=\text{O}$ , or acetone, a compound which has long been known to yield chloroform when distilled with calcium hypochlorite or chlorinated lime. Premising then that the actual process is not described, we may represent the possible production of chloroform from methyl-ketone by the following equation:



It will thus be seen that theoretically every molecule of acetone (58) yields a molecule of chloroform (119.15), in other words, the same amount of chloroform would be yielded by 13½ ozs. of acetone as in the official process is produced from 21 ozs. of pure alcohol. The presence of acetate of calcium as one of the secondary products of the reaction represented above may, at first sight, appear inconsistent with what we know of the oxidation of ketones, but it may be pointed out that Popoff has shown that under the action of strong oxidising agencies they are decomposed, the group  $\text{CO}$  splitting off in combination with one of the alcohol radicals to form an acid; the other radical is set free, in this case, to be converted into chloroform. Whatever may be the exact chemistry of the process there can be no doubt that the product is superior in point of purity to that of the ordinary method. When shaken with sulphuric acid it is not in the slightest degree colored, nor does it leave behind any empyreumatic odor when allowed to evaporate spontaneously from clean paper or from the hand. Its sp. gr. similarly shows it to be pure, and it seems to be peculiarly suited to medicinal use.—"Medicus," in *Jl. and C. Druggist*.

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