

TABLE IV—Continued.

Formula of Reagent.	Molecular Weight.	Name.	Number of Grams in 1 Litre of the Solution.
$(\text{NH}_4)_2\text{MoO}_4$ } HNO_3 }	196	Ammonium Molybdate Solution	75 255 c.c. <i>Con.</i> HNO_3
NH_4Cl	52.4	Ammonium Chloride	105
$\text{K}_4\text{FeC}_6\text{N}_6, 3\text{H}_2\text{O}$	456	Potassium Ferrocyanide	45.6
KSCN	97	Potassium Sulphocyanate	97
$\text{BaCl}_2, 2\text{C}_2\text{O}$	244	Barium Chloride	24.4
$\text{Ba}(\text{NO}_3)_2$	261	Barium Nitrate	26.1
$\text{CaSO}_4, 2\text{H}_2\text{O}$	172	Calcium Sulphate	<i>Saturated: about 2</i>
CaCl_2	111	Calcium Chloride	111
$\text{MgSO}_4, 7\text{H}_2\text{O}$	246	Magnesium Sulphate	246
$\text{MgSO}_4, 7\text{H}_2\text{O}$ } NH_4Cl }		Magnesia Mixture	{ 123 105 }
$\text{FeSO}_4, 7\text{H}_2\text{O}$	278	Ferrous Sulphate	278
$\text{FeCl}_3, 6\text{H}_2\text{O}$	270	Ferric Chloride	270
$\text{Fe}_2(\text{SO}_4)_3, 9\text{H}_2\text{O}$	562	Ferric Sulphate	281
AgNO_3	169	Silver Nitrate	16.9
HgCl_2	271	Mercuric Chloride	54.2
$\text{SnCl}_2, 2\text{H}_2\text{O}$	226	Stannous Chloride	<i>About 113</i>
$\text{PtCl}_4, 8\text{H}_2\text{O}$	480	Platinum Chloride, or Chloroplatinic Acid	96
H_2PtCl_6			
$\text{Co}(\text{NO}_3)_2, 5\text{H}_2\text{O}$	272	Cobaltous Nitrate	27.2
KOH } KI } HgCl_2 }		Nessler's Reagent *	{ 150 62.5 <i>About 31.2</i> }

*In preparing Nessler's reagent, the potassium iodide and the mercuric chloride are to be separately dissolved; the former in 100 c.c., the latter in about 800 c.c. of water, with the aid of heat. When cold, add the second solution to the first, with constant stirring, until a *permanent* precipitate begins to form, then add the potassium hydrate, and make up to 1 litre. After standing for some time, the clear portion of the solution is to be drawn off for use.