

SELECTIONS.

A TRIPLE Bunsen burner is described by Branson in *Chem. Zeitung*. This consists of three Bunsen tubes, so arranged that they may be separated, giving them distinct flames, or brought near together, whereby a large flame is obtained having the qualities of the Bunsen burner.

GUAIACOL. *Tests of Identity and Purity for Guaiacol.*—Marfori, in the *Annal. Chem.*, gives the following reliable reaction for guaiacol: One drop treated with a few of concentrated sulphuric acid gives a purple red color. If traces of cresol (cresol, creosol, &c.) are present the coloration will be of a greyish green. Excepting veratrol and veratrine no other bodies give this reaction.

OIL OF BITTER ALMONDS.—To detect oil of mirbane and oil of bitter almonds warm the suspected oil with black oxide of manganese and sulphuric acid. Nitrobenzol does not lose its odor by this treatment—it rather becomes more pronounced, and after standing for awhile an odor of oil of cinnamon is developed; oil of bitter almonds at first develops a disagreeable odor which, after some time, entirely disappears.

A NEW test for Lead has been discovered by Brenstein (*Pharm. Zeitung*) who found that phosphate of sodium will detect the presence of lead in solution so dilute that ammonia or sulphuric acid will give no precipitate. The solution must first be slightly acidulated with acetic acid, then supersaturated with ammonia, when the addition of phosphate of sodium will cause a white turbidity or precipitate.

BENZOSOL is the name of a benzoic acid ester, prepared from guaiacol, and introduced into materia medica as a substitute for this latter substance in the treatment of phthisis, its claim to superiority resting upon a less disagreeable taste. It occurs in the form of colorless crystals melting at 50° C., nearly insoluble in water, easily soluble in hot alcohol, chloroform, and ether; when pure it is nearly tasteless and odorless.—*Chem. and Druggist*.

FLUORIDE OF SODIUM.—This compound has been experimentally examined by Tappeiner with reference to its physiological action, particularly compared with that of sodic chloride and bromide. He found that it produced effects similar to those of the latter salts, particularly as regards the weariness, inclination to sleep and the retarding of speech and mental processes. On the other hand, he observed that a 0.1 to 5-per-cent. solution has a restraining influence upon the development of bacteria and cocci.

TO RECTIFY TURPENTINE.—A correspondent of the *Microscope* gives this method for use by microscopists:—Take 1 pint of the common turpentine and mix in a quart bottle with 4 fl. oz. of 98-per-cent. alcohol. Agitate well, and let stand until the two fluids separate. Decant the turpentine (which will form the lower layer)

from the alcohol, and mix it with 1 pint of clear water. Agitate thoroughly, and let stand until these two fluids separate, then from the water decant the turpentine (which this time will form the upper layer) and, finally, mix with the turpentine about 1 oz. of powdered starch, and filter through paper.

NON-SETTLING CINCHONA WINE.—The following recipe from the *Pharmaceutische Zeitung* yields a desirable preparation: Fifty parts of cinchona bark are macerated for two days with 50 parts of alcohol and 450 parts of water, the residuum is extracted with the same quantities of alcohol and water, the combined extracts evaporated to 100 parts, filtered after settling, and again evaporated to the consistency of a thin extract. This extract is dissolved with the aid of gentle heat in 50 parts of glycerine, 150 parts of simple syrup, and 1 of tartaric acid, after which 40 parts of cognac and 750 parts of sherry wine are added. The mixture is allowed to stand at least fourteen days in a cool place, and is then filtered. This wine will remain clear if kept in a temperature of about 70° F.

FLUOROFORM. Two parts of iodoform, 2 parts of silver fluoride, and 1 part of chloroform are mixed in a flask, which is cooled to 0°, and is connected with a lead worm kept at 23°, followed by a U-tube containing silver fluoride heated at 100°. The temperature of the flask is allowed to rise gradually, and the gas which is evolved is collected over mercury. It is purified from chloroform vapor by prolonged contact with fragments of dry caoutchouc, and from carbonic oxide by means of a solution of cuprous chloride in hydrochloric acid. The product is a colorless gas with a pleasant odor resembling that of chloroform. It burns with difficulty, with a blue flame and abundant evolution of hydrogen fluoride. It is only slightly soluble in water, chloroform, or benzene, but alcohol dissolves about five times its own volume. Its sp. gr. is 2.445, and it liquefies under a pressure of 40 atmospheres at 20°. If the pressure is suddenly released it solidifies. This gas is fluoroform, CHF_3 . When heated with alcoholic potash at 160° it yields potassium formate and fluoride. Sodium heated in the gas burns brilliantly with disposition of carbon and sodium fluoride and formation of menthane. When silver fluoride and iodoform react in absence of chloroform, much heat is developed, iodine is liberated, and a fluoroiodoform is produced.—*M. Meslans, in Compt. Rend.*

It is recommended in the *Reperioire de Pharmacie* for July 10th, 1890, to dissolve 4 parts of salol in 4 parts of ether, and then add to 30 parts of collodion, and it is stated that the application of the salol-collodion to the affected parts in acute rheumatism will be followed by rapid relief of pain.

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