Sulphate of Strychnia; and secondly, of setting free the chlorine in the hypochlorite solution,—thus enabling free chlorine to unite with the free Strychnia, when the acid separates from the alkaloidal base to unite with the soda in preference. And also there is no doubt, but that less boiling is required when the acid is used, than otherwise, because in the former, the stronger acid will at once displace the weaker, and by this a more soluble salt is formed; whereas it takes long continued boiling to dissolve the very insoluble Strychnate of Strychnia.

The acid also has an effect upon the gummy and starchy principles of the seed, but as to whether the sugar formed, hinders or assists the precipitation of the alkaloid, has not been enquired into; or it may remain perfectly neutral, and this is the most likely.

Where the hypochlorite solution is added to the decoction of the seeds, rendered very acid, an abundant precipitate of a white colour is thrown down, which is a compound of Strychnia and chlorine. What may be its chemical composition, is uncertain, but it cannot be the hydrochlorate because that salt is very soluble; it cannot be free Strychnia, as its reactions given further, are clearly shewn; and if it were hypochlorite, I do not think a similar precipitate would be formed when chlorine water is added to a solution of a salt of Strychnia, and also where a stronger acid is added to the hypochlorite solution, as the sulphuric, it is principally hydrochloric acid, and free chlorine that are given off.

Reasoning from these data, the conclusion arrived at is, that it is some compound of chlorine which is not as yet described, as far as I can find out at least.

Now if we look to some of the other halogen radicals for an answer to the question, we find a compound of iodine, which has not as yet had its analogue in the chlorine series, and this is the ioduretted iodide of an alkali, the most common of which is that of potassium.

I think that the hypochlorite solution which I use is a similar compound in many respects with that of iodine, just mentioned, having sodium as its base. There has been a great deal of argument concerning the substance in question, which I think might be almost settled if it be viewed in this light, and compared with the iodine compound of a similar nature.

From these results, (the effects of combining iodine with different bases) we can expect similar reactions with these bodies, when they are placed in similar positions; and as we find that if iodine be added to iodide of potassium, we have a new compound formed, so are we to expect if chlorine be added to chloride of potassium, or sodium, that we will have an analagous compound of chlorine, which reactions I am not