

## TABLE FOR THE DETECTION OF A SINGLE ACID.

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### I. TO A PORTION OF THE NEUTRAL SOLUTION ADD BARIUM CHLORIDE.

White precipitate, insol. in hydrochloric acid : *silicate, sulphate*.

White precipitate, soluble in hydrochloric acid : *borate, carbonate (effervescence), phosphate, sulphite (odor of sulphur dioxide), tartrate, thiosulphate, (separation of sulphur)*.

Yellow precipitate, insol. in acetic acid : *chromate*.

White precipitate, insol. in acetic acid : *oxalate*.

### II. TO ANOTHER PORTION OF THE SOLUTION ADD SILVER NITRATE.

Precipitate, soluble in dil. nitric acid and in ammonia : white, *borate, carbonate, oxalate, silicate* ; yellow, *arsenite, phosphate* ; red, *arsenate* ; dark red, *chromate*.

Precipitate, insol. in dilute nitric acid, but soluble in ammonia : white, *chloride, cyanide, ferrocyanide (difficultly soluble), hypochlorite, sulphocyanate* ; yellow, *bromide (difficultly soluble)* ; orange red, *ferrocyanide*.

Precipitate, insol. in dilute nitric acid and in ammonia : white, (*ferrocyanide*) ; yellow, (*bromide, iodide*) ; black, *sulphide*, (soluble in concentrated nitric acid).

### III. ACIDS NOT PRECIPITATED BY BARIUM CHLORIDE OR SILVER NITRATE.

Brown ring test (page 124) : *nitrate, nitrite*.

Reaction with sulphuric acid (page 121) : *chlorates*.

\*For the method of bringing the substance to be analyzed into solution, and for preliminary examination, see page 118.