differences between the values of $(P_1 - P_2)/n$ will be somewhat greater. As the $(l-k)$'s in this case, however, are more than twice as great as in the case of thermal expansion, the differences in the values of $(P_1 - P_2)/n$, if expressed as percentages of one of them, will be smaller than in the case of thermal expansion. Accordingly we find from Brückner's observations,

For n	=	0·5	1·0	1·5	2·0	2·5
$(P_1 - P_2)/n$	=	·$0_2$116	$0_2$122	·$0_2$126	·$0_2$128	·$0_2$135

For surface-tension $l-k$ for NaCl is $-0·096$ and for KCl $-0·116$. The approximation to constancy (judged by the percentage criterion) will thus not be so close as in the last case. Rother's observations give, by graphical interpolation,

For n	=	1·0	1·5	2·0
$(P_1 \quad P_2)/n$	=	·016	·0113	·0105

For refractive index $l-k$ for NaCl is $+0·0054$ and for KCl $+0·0091$. Thus the values of $(l-k)\Delta\alpha/\Delta n$ for the two salts are much more nearly equal than in the case of the other properties and consequently the differences in the values of $(P_1 - P_2)/n$ will be smaller than in the case of the other properties. Bender's observations give for the D line,

For n	=	1·0	2·0	3·0
$(P_1 - P_2)/n$	=	·$0_3$29	·$0_3$17	·$0_3$24

If the value for $n=2$ be omitted from consideration, as being probably in error, $(P_1 - P_2)/n$ is seen to be more nearly constant so far as absolute magnitude is concerned than in the other cases considered. As the values of $(P_1 - P_2)/n$ however, are small, their differences when expressed as percentages are comparatively large, and the approximation to constancy, viewed in this way is less than, *e. g.*, in the case of density.

The above account of this phenomenon may be further tested by the aid of Kohlrausch's observations of electrical conductivity; for in this case $l-k$ is the molecular conductivity at infinite dilution (usually written μ_∞). The