

macopcea for 1850. It is prepared by dissolving Strychnia in Mariatic Acid. It is deposited in fine crystals as the concentrated boiling solution cools. They are four-sided needles, which lose their transparency in the air, and when heated give off Hydrochloric Acid. It contains two equivalents of water, and is more soluble than the sulphate. Its uses will be similar to those of pure Strychnine.

CHLORURETTED CHLORIDE OF STRYCHNIA.—This is formed when a solution of chloride of soda, obtained by precipitating the lime from the commercial chloride of that earth, is added to a solution containing Strychnia. It is a white, insoluble precipitate, and is not crystallizable; it is, however, very poisonous.

Strychnia will most likely form salts with the oxygen acids of chlorine, as the hypochlorous, chlorous, chloric, and perchloric acids.

Chlorine, according to Lowig, converts Strychnia into a base containing chlorine, the atoms of hydrogen being replaced by it. He also says that if this haloid element be added to a soluble salt of the alkaloid it causes a white precipitate. This is most likely a similar compound with that formed by chloride of soda, before mentioned.

Silliman, also, states that both Strychnia and Brucia yield products in which the hydrogen is in part replaced by chlorine.

STRYCHNIA AND IODINE—HYDRIODATE OF STRYCHNIA.— St H I or $(\text{C}^{42} \text{H}^{22} \text{N}^2 \text{O}^4) \text{H I}$.—This salt is mentioned by Löwig as being insoluble in water; but he does not describe any other Iodine compound but this. Lewuer also states that a soluble salt of this alkaloid is precipitated white by Iodide of Potassium. Brande states that if Strychnia be boiled in Iodine and water that it is dissolved.

The salt is easily formed by precipitation from Sulphate of Strychnia by Iodide of Potassium. It falls down as a white crystalline powder, which consists of needle-like, irregularly-rounded crystals, and differs from most of the others in not being distinctly cubical. It is not soluble, except in a very great quantity of water, and hence must not be washed too freely in preparation. It does not dissolve in Ammonia, nor in Hydrochloric or Citric Acids, but does in Acetic Acid. Sulphuric and Nitric acids do not act very readily on it if they be dilute; but, when concentrated, the salt is decomposed and Iodine given off. It is freely soluble in alcohol. Its uses will most likely be similar to those of the alkaloid itself. The trivial quantity of Iodine which it contains will have little if any influence when united in small quantities with such a powerful base.

IODURETTED HYDRIODATE OF STRYCHNIA.— St I, I $(\text{C}^{42} \text{H}^{22} \text{N}^2 \text{O}^4)$