titative measurements; and at the close of this research I wish to express my sincere thanks to Prof. W. Lash Miller for his interest and assistance.

The University of Toronto
November 1905.

The experiments were carried out as described in this Journal, **7**, 357, and the abbreviations made use of in these tables are those explained in that paper.

Ac, 10 represents ten cc of 0.050*F* sulphuric acid (0.059 gram-formula-weights H_2SO_4 per liter); *Cr*, 10, ten cc of 0.0083*F* $K_2Cr_2O_7$ (including unreduced bichromate in the "*Ox*"); *KI*, 10, ten cc of 0.0479*F KI*; *Fe*, 1.0, one cc of 0.05*F* ferrous sulphate; *Ox*, 5, the product of oxidation of five cc of 0.05*F* ferrous sulphate; θ is the duration of the reaction in minutes; *As*, the number of cc of *n* 100 sodium arsenite equivalent to the iodine liberated; *V*, the total volume in cc of the reacting mixture.

For the sake of convenience, the tables are numbered consecutively with those of the previous paper, "As, 0" (Table 3)," for example, referring to Table 3, Jour. Phys. Chem. **7**, 3, 9.

I take this opportunity of correcting the following misprints: Table 22 (Jour. Phys. Chem. **7**, 387), second column, should read: 100, 129 152, 162, in place of 100, 116, 126, 131. In equation 7 (p. 372) the sign of equality has been omitted after $dy/d\theta$, and on page 376 the sign has been omitted from the expression $a\ k(1 \quad a) \quad 1077$.

TABLE 24

Ac, 10; Cr, 20; KI, 20; Ox, 5; Fe, 1.0; V, 700

θ	0.5	1	2	4	8
As 30.2 (Expt. 37)		0.55	1.05	1.80	2.40
As, 0 (Table 3)	1.05	1.90	3.10	4.30	5.60

Blanks (30.2 θ 4, As 0; θ 8, As 0.45