University of Toronto.

ANNUAL EXAMINATIONS: 1876.

DEGREE OF M.B.

MEDICINE.

PRACTICAL CHEMISTRY.

Examiner: W. Oliver, B.A.

1. Give a general description of the preliminary process to which you would submit any inorganic substance before making a solution of it.

2. Portions of the same liquid have been tested separately with HCl, H_2S, and AmCl, AmHO, and AmHS, together, giving no precipitate. $AmCO_3$ is then added. What metals may it precipitate, and what leave in solution?

3. Protoxide of mercury is precipitated from a slightly acid solution by a very small quantity of H_2S, but undergoes certain changes on the addition of more of the acid. State these changes, in respect to color particularly.

4. Apply Marsh's test for arsenic to antimony, and state how you would distinguish between the antimony spot on porcelain and the similar one from arsenic.

5. An aqueous or slightly acidulated solution may contain any one of the following metals: silver, mercury, antimony, zinc, calcium; draw up a short table for the purpose of detecting the one present in the solution.

6. A child dies, it is supposed, from having swallowed a small quantity of what is known as "phosphorus paste." The stomach and contents are immediately sent to you for analysis. What would be your mode of procedure? What very delicate test will determine the presence of exceedingly small quantities of this body?