

yellow is complementary to red; violet, being formed of red and blue, is complementary to yellow; and orange, a compound of red and yellow, to blue.

Each of these colors, when placed near its complementary intensifies its effect. Green causes red to appear redder, blue adds a brighter tint to orange, and yellow enhances the color of violet. In other words, every color looks its purest and best when it is beside its complementary.

By means of this simple law of contrast therefore, we have the power of imparting any tint we choose to the complexion, and of rendering our costume agreeable to the eye by the harmony produced by properly contrasted colors; and we may learn to avoid increasing the undesirable tints of a sallow complexion by the juxtaposition of blue or violet, or rendering a florid complexion still more highcolored by contrasting it with green. On the other hand, ladies with golden hair and clear complexions will see that they should wear blue in preference to any other color, harmonizing agreeably, as it does, with the former, without exercising any ill effect on the latter; while those upon whom nature has bestowed dark hair and a rosy complexion should consider green as their especial color.

Very dark people may wear extremely light brown, gray or slate color with impunity, but those who have light brown hair should only venture upon pure white and the light tones of the three primaries, bearing in mind that, where-ever there is agreeable contrasts there is agreeable harmony.—*The Bondior*.

The Monkey and the Drunkard.

Mr. Pollard states that in his drinking days he was the companion of a man in Arundel county, Maryland, who had a monkey that he had valued at a thousand dollars.

We always took him out on chesnut parties, and when he could not shake them off, he would go to the very end of the limbs and knock them off with his fist. One day we stopped at a tavern and drank freely. About half a glass was left, and Jacko drank it up. Soon he was merry, hopped and danced, and set us in a roar of laughter. Jacko was drunk. We all agreed, six of us, that we would come to the tavern next day, and get Jack drunk again, and have sport all day. I called at my friend's house next morning, and we went out for Jack. Instead of being, as usual, on the box, he was not to be seen. We looked inside, and there he was, crouched up in a heap.

"Come out here!" said his master.

Jack came out on three legs; his fore-paw was on his head. Jack had the headache. I knew what was the matter with him; he felt just as I did many a morning. Jack was sick and couldn't go. So we waited three days. We then went, and while drinking, a glass was provided for Jack. But where was he! Skulking behind the chairs.

"Come here Jack, and drink," said his master, holding out the glass to him.

Jack retreated, and as the door was opened, slipped out, and in a moment was at the top of the house. His master went out to call him down, but he would not come. He got a large whip and shook it at him. Jack sat on the ridge-pole and

refused to obey. His master got a gun and pointed it at him. A monkey is much afraid of a gun. Jack slipped over the back side of the house when he saw his predicament, at once whipped upon the chimney, and got down in one of the flues, holding on by his fore-paws. The master was beaten. The man kept the monkey twelve years, but could never persuade him to touch another drop of liquor. The beast had more sense than the man who has an immortal soul, and thinks himself the first and best of God's creatures on earth.

Sulphuretted Hydrogen.

This gas, which for experimental purposes is usually obtained by means of sulphuret of iron, may be procured more conveniently, and in a state of great purity, by the use of sulphuret of calcium. The latter is formed very easily by mixing uncalcined powdered gypsum with one-fourth of its weight of calcined gypsum, and powdered pit-coal equal to one-third of the whole of the gypsum used, and working up the mixture to a stiff dough with water; next forming it into pieces four inches long, two wide, and one and a-half thick, sprinkling them with powdered coal, and drying them, then placing them with coke in a wind furnace, and keeping them at a very high temperature for two hours. When cold they will be found externally to consist of oxysulphuret of calcium; but internally of pure peach-colored sulphurate of calcium, which may be broken in pieces the size of nuts, and preserved in well-stoppered glass bottles. If water is added to these, and then sulphuric acid in small quantities at a time, sulphuretted hydrogen is given off with great uniformity.—*Scientific Review*.

New Test for Acids and Alkalies.

Owing to its property of being reddened by acids and turned blue by alkalies, prepared litmus is perhaps the most generally useful of all the chemical tests which have hitherto been known. A test of a similar nature, but very much more sensitive than litmus, has recently been found by Schönbein to be furnished by the artificial blue coloring matter obtained by acting on chinoline with iodide of amy, and known as "cyanin." This body is so readily acted upon by acids and bases, and its tinctorial power is so enormous, that its delicacy as a test for either acids or bases is quite marvelous. It will detect the presence in water of one-millionth of either sulphuric acid or caustic potash, and of quantities of carbonic acid which cannot be detected by means either of lime or of barytes. Pure distilled water colored with it so as to be quite blue while preserved from contact with the atmosphere has its color instantly destroyed by being blown into from the lungs, by reason of the carbonic acid in the expired breath. Magnesia is incapable of dissolving in water to a sufficient extent to enable the solution to react upon litmus, but pure water in which magnesia has been shaken up gives a most distinct alkaline reaction with cyanin. So does distilled water which has had oxyd of lead shaken up in it, albeit sulphuretted hydrogen, which will detect one part of lead in 350,000 of water, is incapable of showing that any oxyd of lead has been dissolved.—*Mechanics' Mag.*