resting on one of the clinopinacoids. The other angle was measured on a Fuess goniometer. The forms observed were:

$$c = 0P, [001].$$
$$m = \infty P, [110].$$
$$b = \infty P\bar{\infty}, [010].$$

The angles measured were: $m : m = 100° 45'$, and $\angle\beta = 42°$ (nearly). From these data, taking $\bar{b}$ as unity, we have $\dot{a} = 1.238$ (nearly).

The substance sometimes forms crossed twins in which the basal plane is the twinning plane. An orthographic projection of such a twin is shown in Fig. 2.

FIG. 2.

It seems probable that the magnesium chloride, or lithium chloride, present when manganous chloride crystallises out with only two molecules of water must act as a sort of dehydrating agent; and it is not unlikely that the presence of such salts as these in solutions in other cases would lead to the formation of crystals of various substances, containing less than their normal number of molecules of water.

Experiments with Potassium Chloride, etc.

On adding potassium chloride to a hot aqueous solution of manganous chloride, sufficiently concentrated to deposit crystals on cooling, a considerable amount of the alkali chloride is dissolved. When the solution cools crystals of a double chloride are deposited, which are pale pink in color, and usually form radiating groups. The individual crystals are elongated and very thin as produced in this way; but by spontaneous evaporation of the solution they can be obtained in much larger form. They are very soluble in water, but cannot be recrystallised in the ordinary sense of the term, for their solution gives a deposit of potassium chloride only, until the manganous chloride is present in large excess. This decomposition of the salt occurs in the same way in the presence of hydrochloric acid, and seems to depend merely on the different solubilities of the two constituents.

The salt is deliquescent in moist air, and was therefore dried in a desiccator over calcium chloride before analysing. It seems to