

SULPHATES.

(e) 1. Anhydrite Ca SO_4 : 2. Selenite } $\text{Ca SO}_4 + 2\text{H}_2\text{O}$ Gypsum. } Alabaster. } Insoluble in water. May be artificially prepared by precipitating a soluble salt with H_2SO_4 :	Sr SO_4 Celestine. Same.	(e) Ba SO_4 Heavy Spar. Same Blanc Fixe.	(e) Mg $\text{SO}_4 + 7\text{H}_2\text{O}$. <i>Epsom Salts</i> occurring in springs, etc. Is prepared from Dolomite, by separating the lime with sulphuric acid. It forms double sulphates with the alkaline sulphates, one molecule of the water being replaced by the alkaline Sulphate, Mg SO_4, K₂ $\text{SO}_4 + 6\text{H}_2\text{O}$. In this case no precipitate is formed. Mg SO_4 being readily soluble in H_2O.
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NITRATES.

(f) Ca $(\text{NO}_3)_2 + 4\text{H}_2\text{O}$. Formed Ca $\text{CO}_3 + 2\text{HNO}_3 = \text{Ca}(\text{NO}_3)_2 + \text{CO}_2 + \text{H}_2\text{O}$. Soluble in water.	(f) Sr $(\text{NO}_3)_2 + 5\text{H}_2\text{O}$. Same. Same.	(f) Ba $(\text{NO}_3)_2$ Same. Same.	(f) Mg $(\text{NO}_3)_2 + 6\text{H}_2\text{O}$ Same. Same.
(g) Phosphate of Lime $\text{Ca}_3(\text{PO}_4)_2$ exists in bones of animals.			
Fluor Spar Ca F_2 native			
Calcium Sulphide, Ca S, formed in soda ash process.			
Calcium penta-sulphide, CS_5, a soluble compound.			