

V.—HYDROFLUOSILICIC ACID, H_2SiF_6 .

Throws down..

No precipitate, even when an equal volume of alcohol is added.	No prec., even in concentrated solutions. Precipitation only takes place in highly concentrated solutions upon addition of an equal volume of alcohol.	A colored crystalline prec. of Barium Silico Fluoride, $BaSiF_6$ An equal quantity of alcohol hastens and perfects precipitation.	No prec.
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VI.—HYDROGEN SODIUM PHOSPHATE, $HNazPO_4$.

Throws down..

White prec. of $HCaPO_4$ or $Ca_3 2(PO_4)$ —in solutions of Calcium Sulphate, $CaSO_4$, when the phosphate is neutral or alkaline. Soluble in acids, even acetic acid, also in NH_4Cl .	Same as Barium.	A white prec. of Hydrogen Barium Phosphate $HBaPO_4$ out of neutral or alkaline solutions. Free acids dissolve it.	Hydrogen Magnesium Phosphate $HMgPO_4 + 7H_2O$ as a white powder (if Magnesium solutions be not too dilute) which upon boiling becomes $Mg_3(PO_4)_2 + 7H_2O$. If before the addition of the precipitant, Ammonium Chloride and Ammonia be introduced, Ammonium Magnesium Phosphate, $NH_4MgPO_4 + 6H_2O$, as a white crystalline prec., will be thrown down, even in very dilute solutions. This prec. is insoluble in water containing Ammonia, and it is in this form that Magnesium is quantitatively determined.
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