

the powder was in contact with it. It appeared to act independently of absorption.

STRYCHNIA AND BROMINE--**HYDROBROMATE OF STRYCHNIA.**—⁺St HB or $(C^{12}H^{22}N^2O^4)HB$.—I can find no mention made of this salt in the works to which I have access. It can be prepared by precipitating a solution of any salt of Strychnia by the Bromide of Potassium. It is a pretty white salt, which crystallizes in bundles resembling hairs. They are very long and round. It, like the Iodide, is soluble to a slight extent in water; and is so, likewise, in diluted acids. Its uses will, most likely, be similar to those of the alkaloid. Löwig and Silliman say that both Strychnia and Brucia can be converted into bases in which the hydrogen is replaced by Bromine; being, in this respect, similar to chlorine.

BROMIURETTED HYDROBROMATE OF STRYCHNIA.—This is prepared by adding a watery solution of Bromine to a solution of a salt of Strychnia. It is of a yellowish red colour. Its uses will be similar to those of the alkaloid. I gave a kitten $\frac{1}{10}$ th of a grain, which caused sickness for a few hours, but no spasms; but, on this amount being repeated, it proved fatal in about twelve or fifteen minutes, with the usual symptoms.

STRYCHNIA AND FLUORINE.—The Hydrofluorate, although not as yet mentioned by any author, I think could be prepared as most of the homologous compounds are; but I had not the means of examining it.

STRYCHNIA AND CYANOGEN.—The reactions of Cyanogen, although not an elementary body, approach so closely to those of the haloid group, and bear in reality as much relation to them as Acetic, Citric, Tartaric and other vegetable acids to the mineral oxygen acids,—that I think it can with justice be called the organic halogen radical; because in all its combinations it reacts similarly, being only decomposed by great heat or exposure to the air, under certain conditions.

HYDROCYANATE OF STRYCHNIA.—⁺St HCy or $(C^{12}H^{22}N^2O^4)$ (HC²N).—This is a beautiful crystalline salt. Kane says the needles are decomposed by a gentle heat. It is prepared, according to him, by dissolving Strychnia in Hydrocyanic acid; but, as the salt is insoluble in five per cent acid, the mixture does not appear to change, and it is difficult to say when the product is obtained. I think that it is very much better to prepare it by precipitation from a solution of Strychnia by Cyanide of Potassium. The temperature of the solution causes a difference in the appearance. If both solutions be cold, it falls as an amorphous or obscurely crystalline precipitate: under the microscope, it appears like very small needles. If, however, the sulphate of Strychnia