

Pharmaceutical Notes.

DITHIENYL.—Tohl has obtained a body by the oxidation of thiophene with strong sulphuric acid, which he calls dithienyl. It has the composition $C_4N_3S-C_4N_3S$, and is obtained by mixing 5 grams of thiophene and 50 grams of ice-cold water, and shaking this with fuming sulphuric acid. The dithienyl can be separated out by distillation. It crystallizes in large, white plates, and forms a perbromo-derivative, $C_4Br_4S-C_4Br_4S$, which crystallizes in needles, melting at 257° .—*Berichte*.

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THE FRUIT ACIDS AS SOLVENTS FOR QUININE.—Crousee recommends the use of citric and tartaric acids, instead of sulphuric, for dissolving quinine, as the bitter taste is not so strongly developed, and they harmonize better with the digestive fluids, and do not lessen their action. To render soluble 1 gram of quinine sulphate he uses 20 centigrams of tartaric or 60 centigrams of citric acid.

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OIL OF WINTERGREEN.—Bourquelot has just shown that the root of *Polygala vulgaris* contains methyl salicylate. This is another example of the existence of the same body in plants of the same family, for Reuter has already shown its existence in another species of *polygala*.—*Repertoire de Pharmacie*.

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AN EXPLOSIVE MIXTURE.—At the Société de Thérapeutique de Paris, M. Creque called attention to the following dangerous prescription for a tooth powder:

Chlorate of potassium	5 grms.
Borax	10 grms.
Calcined magnesia	10 grms.
Prepared chalk	10 grms.
Oil of peppermint	11 drops.
Saccharine	50 centigrams.

The pharmacist first mixed the chlorate with the saccharine in a mortar, a violent explosion at once resulted, the dispenser was severely burned, and serious damage was caused to the pharmacy, (*Nouv. Remèdes*).—*B. and C. Druggist*.

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MERCURY AND POTASSIUM HYPOSULPHITE.—This is a new agent for hypodermic medication. It presents itself in the form of colorless crystals, easily soluble in water and producing no precipitate in solutions of albumen. The proportion of mercury which it contains is 31.4 to 100. This salt has, according to Dr. Dreser, (*Med. Bull.*), the remarkable property of being decomposed by electrolysis in such a manner that its mercury goes to the anode, for in the preparation the molecule of mercury does not exist in the metallic state, but in the form of a mercuric acid.

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INSOLUBLE GUMS RENDERED SOLUBLE BY HYDROGEN PEROXIDE.—It is stated

(*Bayr. Ind. u. Gewerbebl.*) that if insoluble gums are treated for three or four hours at from 80 to 100° C. with solution of hydrogen peroxide, they will become soluble in water. The proportions given are: Gum, granulated, 220 g., water, 1,000 g.; hydrogen peroxide (12 vol. p. c.) 50 to 60 g. A larger proportion of hydrogen peroxide solution will render the gum completely soluble. Tragacanth is similarly affected.

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CADMIUM SALICYLATE is an energetic antiseptic used in the treatment of purulent ophthalmia and similar affections, whose method of preparation and description are set forth by Cesaris, in the *Boll. chim. farm.* It is made by dissolving cadmium hydrate or carbonate in salicylic acid, or by double decomposition between barium salicylate and cadmium sulphate. If chemically pure, it forms fine white tabular crystals, melting at 300° , and soluble in 24 parts of boiling water, 68 parts at 23° , and in 90 parts at 0° . It is also soluble in alcohol and ether, more freely when warm, but is insoluble in chloroform or benzene. Warm glycerine is likewise a solvent, and from it does not precipitate on cooling.—*B. and C. Druggist*.

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TO MASK THE SMELL OF KEROSENE.—Tempere, in *Corps gras Industrielles*, recommends the addition of amyl acetate to the extent of about 1 per cent. (10 gm. to the litre). This addition, which the editor of the *National Druggist*, has tried, and can recommend, not only modifies the disagreeable odor and converts it into a not unpleasant aromatic, but it seems to make the flame brighter and whiter.—*Nat. Druggist*.

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A REACTION FOR LANOLINE.—Two grammes of slaked lime are mixed with .1 gramme of lanoline and heated. After cooling, the mass is mixed with 5 cc. of water, and the whole shaken with 5 cc. of chloroform. The latter is then poured into an equal volume of concentrated H_2SO_4 , and at the surface is developed a very characteristic red color, due to the cholestérine present.—*Journal de Pharmacie*.

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A TEST FOR VASELINE.—Petroleum jellies are sometimes adulterated with a certain quantity of animal or vegetable fats. These can, of course, easily be detected by the process of saponification, but this is somewhat tedious. A simple test is the following:—Five drops of a saturated solution of potassium permanganate in distilled water are mixed in a mortar with 5 grammes of the jelly. If there is no foreign oil present the color remains. If there is any vegetable or animal oil, the permanganate is reduced, and the color changed to a dirty brown. A roughly quantitative determination can be made if the mass is gradually treated with solution of permanganate until the red color is not discharged.—*Union Pharmaceutique*.

To RENDER vegetable objects translucent Lenz recommends in the *Chemiker Zeitung* a solution of crystallized sodium salicylate in an equal weight of water. Clove oil may also be dissolved in this liquid. Cell membranes are affected less by this agent than by chloral hydrate.

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ZINC SUBGALLATE is described as an odorless, nonhygroscopic, nontoxic, nonirritant, greenish-gray powder, neutral, insoluble in all ordinary liquids, unaffected by light, and containing 41 per cent. of zinc oxide and 56 per cent. gallic acid. It is used both internally and externally.

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ANTIRHEUMATIN.—This body is in the market in the form of dark blue crystals, soluble in alcohol and water, with a slightly bitter taste. It consists of a combination of sodium salicylate and methylene blue. The urine of patients taking it is green, owing to a slight oxidizing action.—*Drog. Zeitung*.

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SALANTOL.—This body is prepared, according to Bourget, by the action of acetone on salicylic acid. It possesses a similar action to that of salol, and is prescribed for summer diarrhoea. It is broken up, in alkaline solutions, into acetone and salicylic acid.—*Rundschau*.

Preparation of Simple Cerate.

In a paper read before the North British Pharmaceutical Society Mr. Boa (*B. and C. Druggist*) said that in the usual method of preparing simple ointments, where the material was constantly stirred while cooling, the effect of such stirring was to incorporate with the ointment a considerable quantity of air. The only reason for stirring he had been able to find in the works referred to was that it was to render the ointment homogeneous. So far as his experiments indicated the reverse was the case. He had prepared simple ointment with excellent results by melting the ingredients together and allowing the product to cool at rest. Ointment so prepared was not lumpy, kept twice as long as the usual product, and possessed other advantages. It was a little harder than the other, the melting point of which was from 1° to 3° lower. If the hardness were an objection, that could be remedied by reducing the quantity of wax. To satisfy himself as to the homogeneity of the ointment, he took pieces from the top, bottom and centre of the ointment, and determined their melting points. In every case they were the same. The advantages gained by letting the ointment cool without stirring were comparative freedom from air and greater keeping power. Other members agreed with him in this matter, and had followed this method for some time.

The successful dealer always finds time to read his trade paper.