needles. These are not soluble in cold water to any appreciable extent, very slightly soluble in boiling water, but easily soluble in alcohol. After one or two recrystallizations, this salt was obtained in a state of purity, and proved to be the hydrochlorate of trimethyl diethyl amido-benzene:

$$C_6 (CH_3)_3 (C_2 H_5)_2 NH_2, HCl,$$

which requires the following values—

	Theory.		Experiment. I.	II.	III.	IV.
C_{13}	67.20	....	66.82	—	—	—
H_{22}	9.65	....	9.72	—	—	—
N	7.47	....	—	7.61	—	—
Cl	15.68	....	—	—	15.6	15.57

The free trimethyl diethyl amido-benzene obtained from this salt by addition of an alkali, is a liquid boiling between 288° and 290°. It has a specific gravity of .971, and is quite colorless when first set free, but darkens and becomes too thick to pour on standing exposed. The hydrochlorate does not yield a well crystallized double platinum salt, but with palladium chloride it forms a beautiful green double salt in feathery crystals.

The acetate and sulphate crystallize in needles, and are very soluble. The oxalate is very difficultly soluble, and crystallizes in prisms.

The Acetyl compound, $C_6 (CH_3)_3 (C_2 H_5)_2 NH (C_2 H_3 O)$, is easily obtained by the action of acetic anhydride; it crystallizes in rosettes of needles, and melts at 182°.

The Isonitril, $C_6 (CH_3)_3 (C_2 H_5)_2 NC$, is best prepared by mixing the base with an alcoholic solution of caustic potash and chloroform in a small flask fitted with a reversed condenser. A violent re-action ensues, and the pungent isonitril odour is quite marked, thus proving the primary character of the base. After neutralizing the excess of base present with sulphuric acid, the isonitril is extracted with ether, which extract yields on evaporation a thick oil which soon crystallises in short prisms, melting between 190 and 192°.